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Synthetically Reversible, Proton-Mediated Nitrite N–O Bond Cleavage at a Dicopper Site

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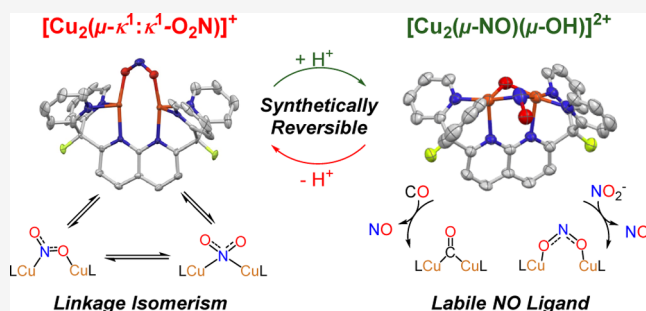
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ABSTRACT: A monocationic dicopper(I,I) nitrite complex $[\text{Cu}_2(\mu\text{-}\kappa^1\text{:}\kappa^1\text{-O}_2\text{N})\text{DPFN}][\text{NTf}_2]$ (**2**) (DPFN = 2,7-bis(fluorodi(2-pyridyl)methyl)-1,8-naphthyridine, $\text{NTf}_2^- = \text{N}(\text{SO}_2\text{CF}_3)_2^-$), was synthesized by treatment of a dicopper acetonitrile complex, $[\text{Cu}_2(\mu\text{-MeCN})\text{DPFN}][\text{NTf}_2]_2$ (**1**), with tetrabutylammonium nitrite ($[\text{nBu}_4\text{N}][\text{NO}_2]$). DFT calculations indicate that **2** is one of three linkage isomers that are close in energy and presumably accessible in solution. Reaction of the $\mu\text{-}\kappa^1\text{:}\kappa^1\text{-O}_2\text{N}$ complex with *p*-TolSH produces nitrous acid (HONO) and the corresponding dicopper thiolate species via an acid–base exchange reaction. Notably, treatment of **2** with HNTf_2 results in N–O bond cleavage in the putative, HONO-ligated complex to form the more thermodynamically favorable nitrosyl-bridged dicopper complex $[\text{Cu}_2(\mu\text{-NO})(\mu\text{-OH})\text{DPFN}][\text{NTf}_2]_2$ (**4**). This scission can be reversed via deprotonation of the hydroxy ligand with KO^tBu . X-ray diffraction studies confirmed the solid-state molecular structures of **2** and **4**. DFT calculations were used to construct a reaction coordinate diagram detailing formation of the $\mu\text{-NO}$ complex and to describe its electronic structure. The nitrosyl ligand in **4** is chemically labile, as demonstrated by its ready displacement in reactions with CO or NO_2^- .



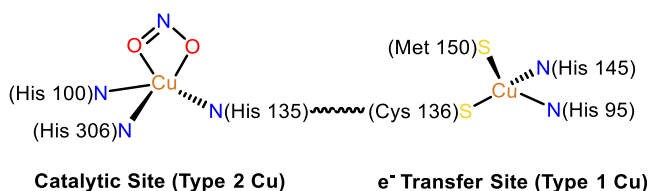
INTRODUCTION

The reduction of nitrite (NO_2^-) to nitric oxide (NO) is a key denitrification step in the nitrogen cycle.^{1,2} The bacterial enzyme copper-containing nitrite reductase (CuNiR) catalyzes this proton-coupled one-electron reaction through the cooperation of type-1 and type-2 Cu sites separated by a Cys-His bridge.^{3–5} Crystallographic studies indicate that in its resting state, the active site (the type-2 Cu ion) is ligated by three histidine residues and a water molecule in a tetrahedral fashion.^{6,7} The type-1 Cu ion is bound by a (His)₂-Cys-Met donor set, consistent with its role as an electron transfer site (Scheme 1a).⁸ The postulated mechanism involves initial displacement of water and binding of a chelating nitrito (O-bound) ligand to the oxidized type-2 Cu site. Protonation by a nearby amino acid residue triggers electron transfer from the type-1 to the type-2 Cu site and subsequent nitrite N–O bond cleavage to produce NO and a Cu(II)–OH species (Scheme 1b).^{9,10} Moreover, the physiological pH determines the efficiency and direction of the reaction: a pH of 6 gives rise to an activity maximum while more basic conditions promote the oxidation of NO back to NO_2^- .^{11,12}

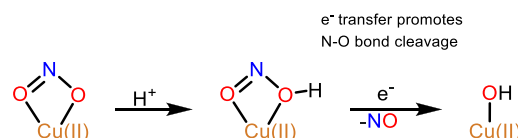
Significant effort has been focused toward the development of molecular copper nitrite complexes that serve as structural and reactive mimics of CuNiR.¹³ Such studies are also expected to aid in the design of catalysts for the removal of nitrogen oxide pollutants from the environment.² These

Scheme 1. (a) CuNiR Active Site with Nitrite Bound; (b) Postulated Mechanism for Nitrite Reduction by CuNiR

a) CuNiR active site



b) Nitrite reduction at CuNiR

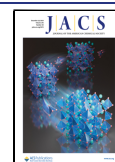


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functional models usually contain a nitrite-bound copper(II) or copper(I) center, which performs the reduction in the presence of a proton source such as acetic acid.^{14–17} Monometallic copper complexes are often the focus of these studies, given that enzymatic catalysis only occurs at the mononuclear, type-2 Cu site. However, the nitrite ligand's ability to engage in several coordination modes and chelate more than one metal center suggests the possibility of multiple activation pathways.¹⁸ Uyeda and Peters have shown that bimetallic Co/Mg sites successfully reduce nitrite, with cobalt acting as the reductant and magnesium as a Lewis acid that aids in activating the N–O bond toward cleavage.¹⁹ Along these lines, viable catalyst design strategies might involve multiple copper centers that cooperate in substrate activation and/or multielectron and multiproton transfer.

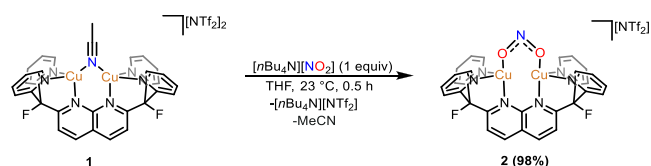
The few reported dicopper nitrite complexes have been primarily investigated for their structural and magnetic properties, rather than for their ability to mediate NO₂[−] reduction.^{18,20–26} Recently, Zhang and co-workers reported the activation of nitrite at a dicopper core to yield a novel [Cu₂(μ-NO)(μ-O)Py₄DMB]²⁺ (Py₄DMB = 1,2-bis(di-pyridine-2-yl)methoxy)benzene) complex that engages in hydrogen atom abstraction (HAA) and C–H bond hydroxylation.²⁷ However, the putative dicopper nitrite intermediate remains undefined, and only the final [Cu₂(μ-O)(μ-NO)]²⁺ product of nitrite activation was observed.

A dicopper platform that permits stable nitrite coordination and selective reduction should contribute to a more comprehensive understanding of the factors controlling N–O bond activation. To this end, this laboratory has employed complexes of the rigid dinucleating ligand 2,7-bis(fluoro-di(2-pyridyl)methyl)-1,8-naphthyridine (DPFN), which supports dicopper complexes that exhibit bimetallic cooperativity and possess unusually high electrophilicity.^{28,29} Herein, we describe a rare dicopper(I,I) nitrite complex that undergoes reversible proton-assisted N–O bond cleavage at the μ-κ¹:κ¹-NO₂[−] ligand, to produce a unique [Cu₂(μ-NO)(μ-OH)DPFN]·[NTf₂]₂ complex.

RESULTS AND DISCUSSION

The acetonitrile ligand of [Cu₂(μ-MeCN)DPFN][NTf₂]₂ (**1**) is readily displaced by nitrite upon reaction with [nBu₄N]·[NO₂] in THF solution, resulting in a color change from orange to dark red. The diamagnetic product [Cu₂(μ-κ¹:κ¹-O₂N)DPFN][NTf₂]₂ (**2**; Scheme 2 and Figure 1), was isolated

Scheme 2. Synthesis of [Cu₂(μ-κ¹:κ¹-O₂N)DPFN][NTf₂]₂ (**2**)



as an analytically pure brown powder in 98% yield by precipitation from the reaction solution upon addition of diethyl ether. The infrared spectrum of **2** contains a sharp band at 1406 cm^{−1} that shifts to 1378 cm^{−1} with ¹⁵N labeling at the μ-κ¹:κ¹-NO₂[−] ligand (Δν(NO₂)_{14N–15N} = 28 cm^{−1}, Figure S24). The ¹⁵N NMR spectrum of the enriched sample possesses a single resonance at 556.5 ppm, which is modestly

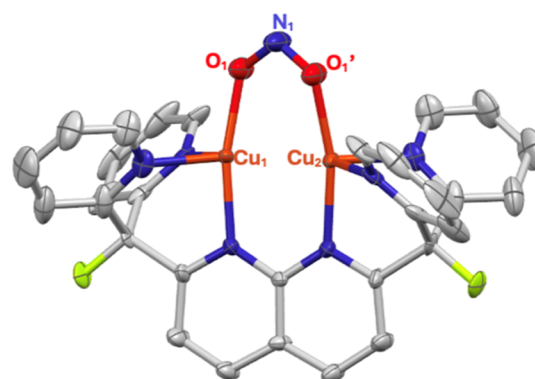


Figure 1. Solid-state structure of **2** as determined by single-crystal X-ray diffraction. The NTf₂[−] counterion and all hydrogens are omitted for clarity. Thermal ellipsoids are set at the 50% probability level.

shifted with respect to the resonance of NaNO₂ in water (δ = 609.5 vs liquid NH₃).

The solid-state molecular structure of **2**, determined by single crystal X-ray diffraction, revealed that the copper(I) ions exist in an unusual, distorted seesaw geometry (τ₄ ≈ 0.65, τ₄' ≈ 0.44)^{30,31} with an average Cu–Cu distance of 2.55(3) Å. The nitrite bonding metrics in the three cationic complexes within the asymmetric unit cell exhibit statistically significant variations, likely resulting from molecular motion. Nevertheless, the symmetrically O,O-bound nitrito ligand (Figure 1) displays surprisingly short Cu–O contacts (Cu–O_(avg) = 1.96(2) Å), indicating strong bonding interactions. The enhanced bonding interactions may also result from the reduced ring strain inherent to a μ-κ¹:κ¹ binding mode as compared to a nitrite ligand chelating a single copper. The N–O bond distances (N–O_(avg) = 1.23(11) Å) are comparable to those found in the few mononuclear copper(I) nitrito complexes that have been reported.^{32–37} To our knowledge, **2** is the only dicopper(I,I) μ-κ¹:κ¹-nitrite complex to be structurally characterized. In accordance with hard–soft acid–base (HSAB) theory, the nitrito binding mode is most common in cupric systems.^{38,39} Consequently, mononuclear copper(I) nitrito complexes are often unstable.^{40,41} The few that have been reported are almost exclusively stabilized by soft phosphine ligands, with only one possessing a bidentate N-donor ligand.^{32–37} In contrast, complex **2** is stable indefinitely in the solid state and in solution. A possible explanation for the enhanced stability of **2** is that bonding to two copper(I) ions results in a coordinatively nonlabile nitrito ligand.

While multinuclear NMR spectroscopy and X-ray crystallography indicate a symmetrical μ-κ¹:κ¹-nitrite binding mode, DFT computations (B3LYP/def2-TZVPP/CPCM-THF) suggest that the ligand can engage in several coordination modes akin to those of previously reported mononuclear copper nitrite complexes.^{42–44} Two closed-shell singlet linkage isomers were found to be close in free energy to the O,O μ-κ¹:κ¹ model complex (**2***). The model linkage isomer closest in energy (+0.5 kcal mol^{−1}) to the observed ground state features a μ-(η¹-N:η¹-O) (**2*'**) binding mode for the bridging nitrite (Figure 2a), reminiscent of the nitrite-bridged cobalt–magnesium complex reported by Uyeda and Peters.¹⁹ The other linkage isomer, with a nitro (N-bound) binding mode (**2*''**), is only 1.6 kcal mol^{−1} higher in free energy than the ground state (Figure 2b). Calculations indicate that higher solvent dielectric constants lessen the energy gap between the isomers (Figure S34). A small energetic difference between

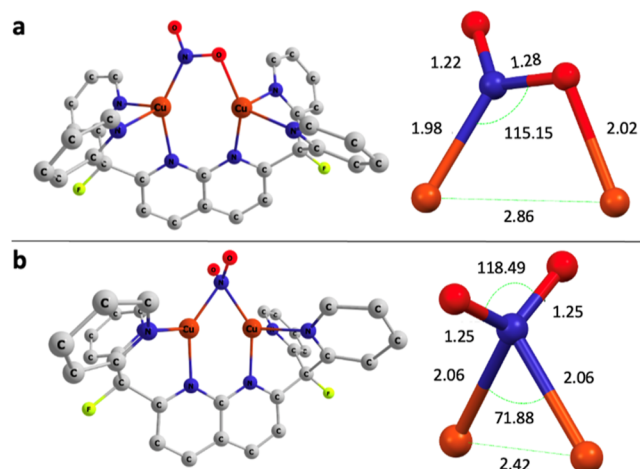
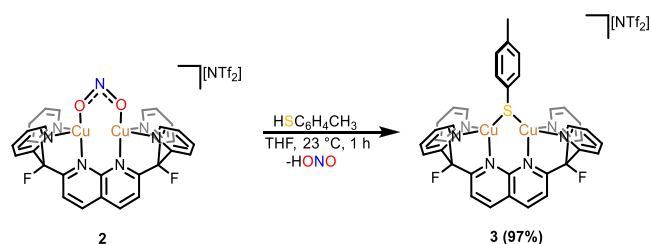


Figure 2. DFT calculated molecular structures of closed-shell singlet linkage isomers of $[\text{Cu}_2(\mu\text{-}\kappa^1\text{:}\kappa^1\text{-O}_2\text{N})\text{DPFN}]^+$ (2^*). The $[\text{Cu}_2(\mu\text{-}\eta^1\text{:}\eta^1\text{-NO}_2)\text{DPFN}]^+$ ($2^{*'}\text{a}$) and the $[\text{Cu}_2(\mu\text{-}\eta^1\text{-NO}_2)\text{DPFN}]^+$ ($2^{*'}\text{b}$) linkage isomers are 0.5 and 1.6 kcal mol⁻¹ higher in free energy than the ground state, respectively. All hydrogens are omitted for clarity.

linkage isomers is not unprecedented for copper(I) nitrite complexes,^{41–45} and in this case the gap is likely minimized by the presence of a second Lewis acidic center that aids in accommodating different binding modes. These computational results indicate that all isomers are energetically accessible for reaction chemistry.

To ascertain its capability for proton-coupled nitrite reduction, complex **2** was treated with acids of varying strength. Thiols were of particular interest given their ubiquity in metabolic processes and the reported protonation of a nonheme dinitro complex to produce NO.^{46–48} Treatment of **2** with one equiv of *p*-TolSH in THF solution resulted in a rapid solution color change from dark red to orange-red. The ¹H NMR spectrum of the crude mixture indicated consumption of the nitrite complex and appearance of new resonances for $[\text{Cu}_2(\mu\text{-STol})\text{DPFN}][\text{NTf}_2]$ (**3**; Scheme 3). The product was isolated after workup as an analytically pure beige powder in high yield (97%).

Scheme 3. Synthesis of $[\text{Cu}_2(\mu\text{-STol})\text{DPFN}][\text{NTf}_2]$ (**3**)



Complex **3** is likely formed by an acid–base exchange akin to the reactivity previously observed for a tris-pyrazolyl zinc nitrite complex reported by Warren and co-workers.⁴⁹ Although the expected nitrous acid byproduct was not directly observed, exposure of cobalt(II) 5,10,15,20-tetraphenylporphyrin (CoTPP) to the gaseous product released from the reaction produced the $\{\text{CoNO}\}^8$ product (61%) and water, as determined by ¹H NMR spectroscopy (Figure S11). The formation of these byproducts in the reaction vessel is likely a result of nitrous acid decomposition after expulsion from the

dicopper core.¹³ The solid-state molecular structure of **3** displays a thiolate ligand bridging the dicopper unit with the tolyl substituent canted away from the plane containing the naphthyridine backbone (Figure 3).

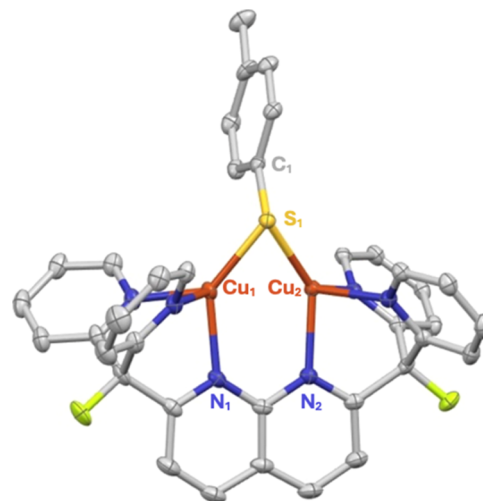


Figure 3. Solid-state structure of **3** as determined by single-crystal X-ray diffraction. The NTf_2^- counterion and all hydrogens are omitted for clarity. Thermal ellipsoids are set at the 50% probability level.

Since the reaction of **2** with a thiol results in displacement of the presumed nitrous acid product formed by protonation, a Brønsted acid with a weakly coordinating conjugate base was employed to favor HONO retention at the dicopper core. Treatment of **2** with HNTf_2 in thawing THF at -196°C , followed by warming to 23°C , resulted in a color change from dark red to dark green, corresponding to the formation of $[\text{Cu}_2(\mu\text{-NO})(\mu\text{-OH})\text{DPFN}][\text{NTf}_2]_2$ (**4**; Figures 4 and 5).

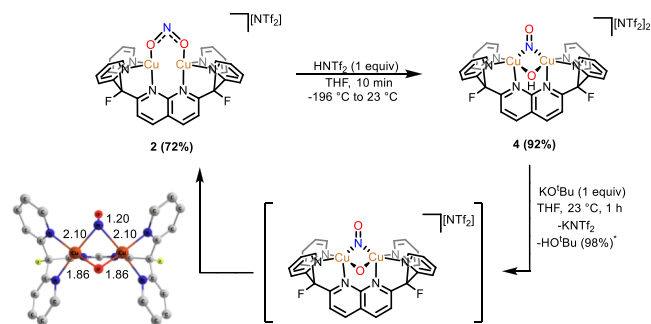


Figure 4. Stoichiometric cycle illustrating the synthesis of $[\text{Cu}_2(\mu\text{-NO})(\mu\text{-OH})\text{DPFN}][\text{NTf}_2]_2$ (**4**) and its deprotonation to yield $[\text{Cu}_2(\mu\text{-}\kappa^1\text{:}\kappa^1\text{-O}_2\text{N})\text{DPFN}][\text{NTf}_2]$ (**2**) via a dicopper(II,II) $\mu\text{-NO}$, $\mu\text{-O}$ intermediate that undergoes N–O bond formation. The DFT-calculated geometry of $[\text{Cu}_2(\mu\text{-NO})(\mu\text{-O})\text{DPFN}]^+$ is also shown with hydrogens omitted for clarity. *Yield determined by integration of the ¹H NMR spectrum against an internal standard.

Addition of pentane to a saturated THF solution of the compound enabled its isolation as an analytically pure dark green powder in high yield (92%). The diamagnetic complex exhibits a ¹H NMR spectrum reflecting a loss of symmetry evident by the presence of 11 resonances that together integrate to 21 protons. The ¹⁹F{¹H} NMR spectrum reveals a shift of the ligand resonance from -172.90 ppm in **2** to -174.41 ppm (with respect to the resonance of benzene-*f*₆ in

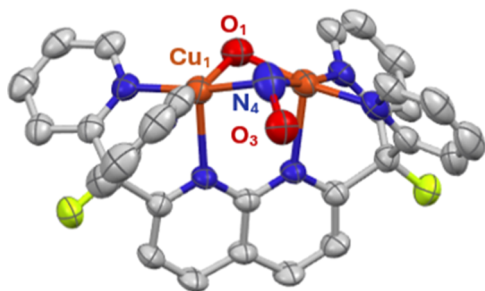


Figure 5. Solid-state structure of **4** as determined by single-crystal X-ray diffraction. The NTf_2^- counterions and hydrogens are omitted for clarity. Thermal ellipsoids are set at the 50% probability level.

THF) and the presence of two NTf_2^- counterions per formula unit (Figure S9). Infrared spectroscopy confirms the presence of two new bands belonging to **4**, a broad band at 3547 cm^{-1} , assigned as an O–H stretch, and another weak band at 1555 cm^{-1} . The latter band shifts to 1522 cm^{-1} with ^{15}N labeling ($\Delta\nu(\text{NO})_{^{14}\text{N}-^{15}\text{N}} = 33\text{ cm}^{-1}$, Figure S28) and is assigned as an N–O stretch. This low $\nu\text{N}-\text{O}$ value is consistent with those reported in the literature for a NO^- ligand bridging an oxidized dicopper core.^{27,50,51}

The obtained solid state molecular structure for this compound is a mixture (30:70) of **4** and the previously reported $[\text{Cu}_2(\mu\text{-OH})_2\text{DPFN}][\text{NTf}_2]_2$ (**5**), likely resulting from unavoidable exposure of single crystals of **4** to air during the diffraction experiment.⁵³ Further collection attempts yielded X-ray diffraction data with similar mixtures. While the structural data illustrates the NO binding mode and the lack of NO/OH ligand disorder across the plane containing the naphthyridine backbone, the structural metrics are less informative as they represent a superposition of both species. Nonetheless, both copper atoms have a slightly distorted square pyramidal geometry ($\tau_5 \approx 0.04$)⁵² and are $2.7569(3)\text{ \AA}$ apart. This distance is much larger than that of **2**, which suggests that the nitrite functions to bring the copper ions closer together. The hydroxy ligand symmetrically bridges the dicopper core ($\text{Cu}-\text{O} = 1.9047(11)\text{ \AA}$) in a manner similar to that of other hydroxy-bridged dicopper complexes supported by DPFN ($1.89\text{--}1.96\text{ \AA}$).^{53,54} The nitrosyl N–O bond distance ($\text{N}-\text{O} = 1.3816(14)\text{ \AA}$) is longer than those of the other reported dicopper μ -nitrosyl complexes, with the Zhang group's $[\text{Cu}_2(\mu\text{-NO})(\mu\text{-OMe})\text{Py}_4\text{DMB}]^{2+}$ complex being the closest at $1.32(5)\text{ \AA}$. The $\text{Cu}-\text{N}_\mu$ bond distance [$1.9601(16)\text{ \AA}$] in the complex is the smallest reported for such a moiety, but remains consistent with structural data of the other nitrosyl-bridged dicopper complexes.^{27,50,51,55}

Computational modeling shows that the ground state of **4** is best described as a dicopper(II,II) open-shell singlet. The copper(II) ions show no coupling between each other, but instead engage in antiferromagnetic interactions with the nitrosyl ligand. This is in good agreement with the diamagnetic character reflected in sharp NMR resonances of the complex. Analysis of its Löwdin spin population shows *ca.* $+0.4$ unpaired e^- on each of the copper atoms. The remaining spin density is mostly located on the nitrosyl ligand with approximately -0.6 unpaired e^- on each of its atoms. The oxygen atom on the hydroxy ligand minimally contributes to the antiferromagnetic behavior with a spin density of $+0.1\text{ }e^-$. Together, the computational and spectroscopic results indicate that complex **4** is formally a dicopper(II,II) complex possessing one OH^- ,

one NO^- , and two NTf_2^- counterions completing the charge balance.

DFT computations (B3LYP/def2-TZVPP/CPCM-THF) were also employed to gain insight into the mechanism of protonation and nitrite cleavage from the model complex $[\text{Cu}_2(\mu\text{-}\kappa^1\text{-}\kappa^1\text{-O}_2\text{N})\text{DPFN}]^+$ (**2***) to yield $[\text{Cu}_2(\mu\text{-NO})(\mu\text{-OH})\text{DPFN}]^{2+}$ (**4***). The calculations suggest that protonation initially occurs at one of the oxygen atoms of the $\mu\text{-}\kappa^1\text{-}\kappa^1$ bridging nitrite, which requires 8.4 kcal mol^{-1} of free energy and forms a transient $[\text{Cu}_2(\mu\text{-}\kappa^1\text{-}\kappa^1\text{-HONO})\text{DPFN}]^{2+}$ complex, intermediate **A**, analogous to those spectroscopically observed during nitrite reduction by functional mimics of the CuNiR.⁵⁶ The following step involves an exergonic linkage isomerism of the HONO ligand to intermediate **B** with a nitro binding mode ($\mu\text{-NO}_2\text{H}$). This is in good agreement with calculations showing that **2*** possesses a more negative free energy of protonation compared to the other two linkage isomers of the nitrite complex ($-2.46\text{ kcal mol}^{-1}$ for **2*** vs $8.43\text{ kcal mol}^{-1}$ for **2***, Figure S35). Intermediate **B** then undergoes an oxidative addition of the N–OH bond at the dicopper core to yield **4***. Cleavage of the N–O bond of intermediate **B** is accompanied by a moderate release of energy ($11.1\text{ kcal mol}^{-1}$) which drives the process forward (Figure 6). Another

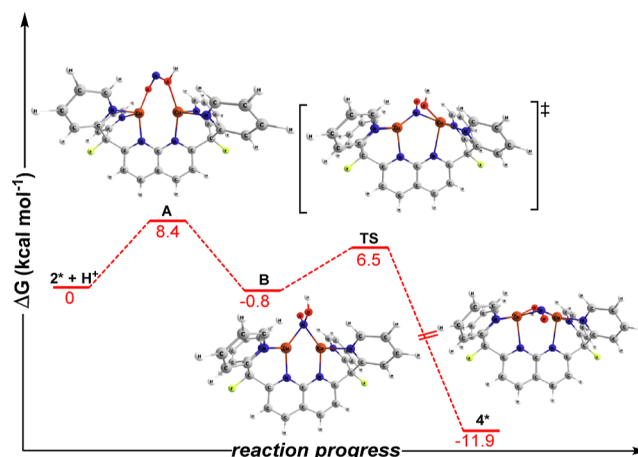


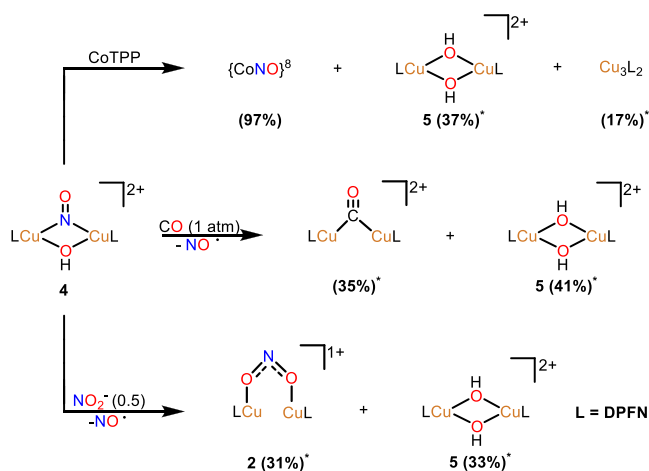
Figure 6. Reaction coordinate diagram for the formation of **4*** derived by DFT at the B3LYP/def2-TZVPP/CPCM-THF level of theory.

possible reaction pathway involves cleaving one N–O bond in $\mu\text{-}\kappa^1\text{-}\kappa^1\text{-NO}_2^-$ prior to protonation. However, to form a $[\text{Cu}_2(\mu\text{-NO})(\mu\text{-O})\text{DPFN}]^+$ species from complex **2***, the system must overcome a significant energetic barrier of 45 kcal mol^{-1} (Figure S36). Similar calculations conducted on the CuNiR active site have attributed the large energetic cost to the destabilizing charge that develops on the O^{2-} atom as the N–O bond of the nitrite ligand is being cleaved.⁹ Thus, reminiscent of the CuNiR mechanism, the protonation-induced N–O cleavage pathway is more accessible.

Nitrosyl complex **4** bears an unusual level of thermal stability given that, to our knowledge, only one other reported dicopper nitrosyl-bridged compound is stable at ambient temperature.⁵¹ Under an inert atmosphere, complex **4** is stable for several days in THF solution or over months in the solid state without signs of decomposition. Interestingly, DFT analysis reveals that there is a $18.9\text{ kcal mol}^{-1}$ barrier to the thermoneutral dissociation of NO from **4** (ΔG of $-0.25\text{ kcal mol}^{-1}$, Figure S38). Treatment of **4** with stoichiometric amounts of CoTPP

resulted in immediate formation of the $\{\text{CoNO}\}^8$ species (97%) along with $[\text{Cu}_3(\text{DPFN})_2][\text{NTf}_2]_3$ (17%) and $[\text{Cu}_2(\mu\text{-OH})_2\text{DPFN}][\text{NTf}_2]_2$ (37%) (**5**), as determined by ^1H and $^{19}\text{F}\{^1\text{H}\}$ NMR spectroscopy. Both copper complexes have been isolated and fully characterized in this laboratory, and the tricationic helical complex is known to result from loss of the bridging ligand at the dicopper core.^{53,57} This indicates a dissociative loss of the nitrosyl ligand from the dicopper center, followed by reaction with CoTPP to yield $\{\text{CoNO}\}^8$. Nonetheless, a bimolecular reaction between both metal complexes resulting in NO transfer cannot be completely ruled out. A proposed pathway involves an initial trapping of NO by Co(TPP), followed by disproportionation of the resulting $[\text{Cu}^{\text{I}}\text{Cu}^{\text{II}}(\mu\text{-OH})]^{2+}$ core to form the stable complexes $[\text{Cu}^{\text{II}}_2(\mu\text{-OH})_2]^{2+}$ and $[\text{Cu}_3(\text{DPFN})_2][\text{NTf}_2]_3$ (Scheme 4).

Scheme 4. Reactivity of $[\text{Cu}_2(\mu\text{-NO})(\mu\text{-OH})\text{DPFN}][\text{NTf}_2]_2$ (**4**) with CoTPP, CO, and $[n\text{Bu}_4\text{N}][\text{NO}_2]^{a}$



^a*Yield determined by integration of the ^{19}F NMR spectrum against an internal standard.

To further probe the lability of the nitrosyl ligand, **4** was exposed to excess carbon monoxide, leading to a gradual color change from dark green to light blue over the course of 30 min (Figure S18). The ^{19}F NMR spectrum of the crude mixture shows the formation of $[\text{Cu}_2(\mu\text{-CO})\text{DPFN}][\text{NTf}_2]_2$ and **5** in 35 and 41% yield, respectively (Scheme 4).⁵⁴ These results suggest that, while CO can competently displace NO, the proposed $[\text{Cu}^{\text{I}}\text{Cu}^{\text{II}}(\mu\text{-OH})(\mu\text{-CO})\text{DPFN}]^{2+}$ species is susceptible to disproportionation. Interestingly, CO is not the only ligand that promotes this type of disproportionation. Treatment of **4** with half an equivalent of $[n\text{Bu}_4\text{N}][\text{NO}_2]$ results in a darkening of the solution to dark red in under 5 min. Multinuclear NMR spectroscopic analysis reveals the formation of **2** (31%), **5** (33%), and the salt metathesis product $[n\text{Bu}_4\text{N}][\text{NTf}_2]$ (95%) (Scheme 4). This reaction also likely proceeds through a mixed-valent intermediate $[\text{Cu}^{\text{I}}\text{Cu}^{\text{II}}(\mu\text{-OH})(\mu\text{-}\kappa^1\text{:}\kappa^1\text{-NO}_2)\text{DPFN}]^+$, which could further react to ultimately yield **2**. Not only do these results establish novel reactivity pathways for a μ -nitrosyl, μ -hydroxy dicopper core, they are also of potential relevance to current interest in nitric oxide releasing systems.⁵⁸

Given the observed stability of $\text{Cu}_2(\mu\text{-}\kappa^1\text{:}\kappa^1\text{-O}_2\text{N})\text{DPFN}][\text{NTf}_2]_2$, it was hypothesized that deprotonation of the hydroxy ligand of **4** would result in N–O bond formation and reductive

elimination of NO_2^- . Indeed, addition of a stoichiometric amount of KO^tBu to a THF solution of **4** gave an instant color change to dark red along with a noticeable exotherm evidenced by warming of the reaction flask. The ^1H and ^{19}F NMR spectra of the crude reaction mixture confirmed successful deprotonation of the hydroxy ligand of **4** to yield **2** (72%), *tert*-butanol (98%), and potassium bistriflimide (Figure 4). Computational modeling (B3LYP/def2-SVP/CPCM-THF) of a likely initial product of deprotonation $[\text{Cu}_2(\mu\text{-NO})(\mu\text{-O})\text{DPFN}][\text{NTf}_2]$, suggests a similar electronic structure to **4**. The optimized structure of this intermediate shares similar Cu–Cu (2.756 Å) and Cu(II)–O (1.859 Å) distances to those observed in $[\text{Cu}_2(\mu\text{-NO})(\mu\text{-O})\text{Py}_4\text{DMB}]^{2+}$. The Cu– N_μ distances (2.104 Å), however, are longer than that of $[\text{Cu}_2(\mu\text{-NO})(\mu\text{-O})\text{Py}_4\text{DMB}]^{2+}$ (Cu(II)– $N_\mu = 2.03(1)$ Å) and those in the nitrosyl-bridged dicopper complex (Figure 4). This intermediate is significantly destabilized relative to nitrite-bridged dicopper complex **2** by 29.6 kcal mol⁻¹ and overcomes a 13.9 kcal mol⁻¹ barrier to N–O bond formation (Figure S37).

CONCLUSIONS

Mononuclear copper coordination compounds have been extensively studied in the context of biomimetic nitrite reduction. Meanwhile, nitrite bond activation at dicopper cores has only recently garnered attention, despite the valuable information these systems can offer. These results demonstrate that two copper(I) centers in proximity can work in concert to stabilize what historically has been an unstable linkage for mononuclear complexes. Moreover, this cooperativity allows the nitrito ligand to be reversibly cleaved by a proton, a process permitted by nitrosyl retention at the dicopper core. To our knowledge, this is the first instance of a stoichiometric cycle reminiscent of the CuNiR's pH-dependent, bidirectional catalysis occurring on a dicopper platform. Computational and spectroscopic methods were used to describe a rare antiferromagnetically coupled dicopper(II,II) nitrosyl hydroxy complex that possesses unusual thermal stability. The lability of the nitrosyl ligand provides a new avenue for the formation and study of other reactive bimetallic species and possesses relevance toward the study of NO-releasing systems.

ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge at <https://pubs.acs.org/doi/10.1021/jacs.4c14642>.

Detailed experimental methods, including characterization data, spectra, details of X-ray crystallography, and computational studies (PDF)

Accession Codes

Deposition numbers 2391275–2391277 contain the supplementary crystallographic data for this paper. These data can be obtained free of charge via the joint Cambridge Crystallographic Data Centre (CCDC) and Fachinformationszentrum Karlsruhe Access Structures service.

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Notes

The authors declare no competing financial interest.

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