Title

Resolving the Model-Observation Discrepancy in the Mesospheric and Stratospheric HOx Chemistry

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Key Points:
- The kinetic rate coefficients and photo absorption coefficients are adjusted to better simulate the observed middle atmospheric OH and HO_2
- The model-observation discrepancy is reduced by adjusting the rate of $H + O_2 + M \rightarrow HO_2 + M$ and the O_2 absorption cross section at Lyman α
- A hitherto unsuspected radiative association reaction, $H + O_2 \rightarrow HO_2 + h\nu$, may play a significant role in the mesospheric HO_x chemistry

Supporting Information:
- Supporting Information S1

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Resolving the Model-Observation Discrepancy in the Mesospheric and Stratospheric HO_x Chemistry

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Abstract We examine the middle atmospheric odd-hydrogen (HO_x) chemistry by comparing the Aura Microwave Limb Sounder (MLS) OH and HO_2 measurements with a photochemical model simulation. The model underestimates mesospheric OH and HO_2 concentrations if the standard chemical kinetic rates are used, whether the model H_2O and O_3 are constrained with observations or not. To resolve the discrepancies, we adjust the kinetic rate coefficients of three key reactions ($O + OH \rightarrow O_2 + H$, $OH + HO_2 \rightarrow H_2O + O_2$, and $H + O_2 + M \rightarrow HO_2 + M$) and the O_2 photo absorption cross section at Lyman α (121.57 nm) using the Bayesian optimal estimation. A much better model-observation agreement can be achieved if the kinetic rate coefficients for $H + O_2 + M \rightarrow HO_2 + M$ is increased by 134–310%, and the O_2 photo absorption cross section at Lyman α is reduced by 33–54%, while the kinetic rate coefficients for $O + OH \rightarrow O_2 + H$ and $OH + HO_2 \rightarrow H_2O + O_2$ remain consistent with the current laboratory values. The kinetic rate coefficient for $H + O_2 + M \rightarrow HO_2 + M$ requires a very large adjustment beyond the uncertainty limits recommended in the NASA Data Evaluation, suggesting the need for future laboratory measurements. An alternative explanation is that the radiative association reaction, $H + O_2 \rightarrow HO_2 + h\nu$, plays a significant role, which has never been measured. Our results demonstrate that high-quality satellite observations can be used to constrain photochemical parameters and help improve our understanding of atmospheric chemistry.

1. Introduction

Odd hydrogen (HO_x) species, including hydroxyl radical (OH) and hydroperoxyl (HO_2), are important catalysts of odd oxygen in the middle atmosphere (Brasseur & Solomon, 2005). The main source of middle atmospheric HO_x is direct photolysis of H_2O by the solar Lyman α line in the mesospheric region (>60 km):

$$H_2O + h\nu\ (121.57\ \text{nm}) \rightarrow H + OH \quad \text{(R1)}$$

or the photolysis of O_3 and N_2O by solar UV below 200 nm and 330 nm, respectively, in the stratospheric region (<60 km) that produces $O(^1D)$:

$$O_3 + h\nu\ (< 200\ \text{nm}) \rightarrow O(^1D) + O_2, \quad \text{(R2)}$$

$$N_2O + h\nu\ (< 200\ \text{nm}) \rightarrow O(^1D) + N_2, \quad \text{(R3)}$$

followed by

$$H_2O + O(^1D) \rightarrow 2OH, \quad \text{(R4)}$$

$$CH_4 + O(^1D) \rightarrow CH_3 + OH, \quad \text{(R5)}$$

$$H_2 + O(^1D) \rightarrow H + OH. \quad \text{(R6)}$$

OH is then converted to HO_2, and vice versa, via reactions with O, O_3, and NO:

$$OH + O_3 \rightarrow HO_2 + O_2, \quad \text{(R7)}$$

$$HO_2 + O \rightarrow OH + O_2, \quad \text{(R8)}$$

$$HO_2 + NO \rightarrow OH + NO_2. \quad \text{(R9)}$$

Throughout the whole middle atmosphere, the ultimate sink of HO_x is through

$$OH + HO_2 \rightarrow H_2O + O_2 \quad \text{(R10)}$$

(Brasseur & Solomon, 2005; Canty & Minschwaner, 2002). Reactions (R1)–(R3) show that the net source of HO_x is sensitive to variations of incoming solar UV solar spectral irradiance (SSI). Satellite observations suggest that the HO_x species better correlate with SSI than O_3 or temperature (Rozanov et al., 2006) and are good indicators of solar cycle with almost zero time lag (Shapiro et al., 2012; Wang et al., 2015).

The HO_x profiles have been observed using balloon-based measurements (Englert et al., 2000; Heaps & McGee, 1985; Jucks et al., 1998; Kendall & Clark, 1980; Park & Carli, 1991; Pickett & Peterson, 1993; Traub et al., 1990), ground-based measurements (Burnett & Burnett, 1981; Cageao et al., 2001; Iwagami et al., 1995), and rocket-borne (Anderson, 1971) and space-borne measurements (Conway et al., 1999; Pickett, 2006). Despite the above simple HO_x photochemistry in the stratosphere and mesosphere, a number of studies reveal discrepancies between observed and simulated HO_x concentrations. Conway et al. (2000) first claimed that the simulated stratospheric OH is lower than that observed by the Middle Atmosphere High Resolution Spectrograph Investigation (MAHRSI) (Summers et al., 1997), while the simulated mesospheric OH is higher. They thus coined the term "HO_x dilemma" to describe this discrepancy having opposite signs in the stratosphere and mesosphere. Later, this apparent dilemma has been attributed to large uncertainties of MAHRSI data at low altitudes (Englert et al., 2008). However, other model-observation discrepancies persist. For example, Millán et al. (2015) showed that the standard photochemistry significantly underestimates the mesospheric HO_2 at 70 km observed by Microwave Limb Sounder (MLS) (Livesey et al., 2015; Pickett et al., 2006; Pickett et al., 2008) and the HO_2 discrepancy disappeared only when the observed OH was used to constrain the model. One possible cause of the aforementioned model-observation discrepancy in the HO_x concentrations may be model biases due to laboratory uncertainties of chemical kinetics rates (Sander et al., 2011). A number of groups (e.g., Canty et al., 2006; Conway et al., 2000; Jucks et al., 1998; Siskind et al., 2013; Summers et al., 1997) tried to adjust the kinetic rates of some important photochemical reactions to better fit the simulated HO_x with the observations. However, the choices of the photochemical reactions to be adjusted are not unique. Table 1 lists the reactions adjusted by some of the previous groups. For example, Canty et al. (2006) adjusted the reaction rates for $OH + HO_2 \rightarrow H_2O + O_2$ (R10) and

$$O + OH \rightarrow O_2 + H \quad \text{(R11)}$$

while Siskind et al. (2013) adjusted the reaction rate for

$$H + O_2 + M \rightarrow HO_2 + M. \quad \text{(R12)}$$

The adjustments of the reaction rates also vary significantly among different studies. To have a more objective choice of reactions to be adjusted, we shall adopt a Bayesian optimal estimation approach

Table 1
Perturbations (in %) of Kinetic Rate Coefficients in Previous Studies and the Current Study

	Summers97[a]		Jucks98[b]		Conway00[c]			This work[f]			
	B	C	D	D1[k]	B[l]	Canty06[d]	Siskind13[e]	I[n]	II[n]	III[n]	IV[o]
$H_2O + O(^1D)$[g]				+25							
$OH + O$[h]				+25	(–50)	+20[m]		(–53)	(–31)	(–15)	+12
$HO_2 + O$	(–50)	(–20)	(–25)								
$OH + HO_2$		+30	(–25)			+20		+30	(–13)	(–15)	(–10)
$H + O_2 + M$[i]							+52[p]		+131	+134	+310
$O_2 + h\nu$[j]										(–54)	(–33)
Ly α							+8				

Note. The previous studies included here are Summers et al. (1997) (Summers97), Jucks et al. (1998) (Jucks98), Conway et al. (2000) (Conway00), Canty et al. (2006) (Canty06), and Siskind et al. (2013) (Siskind13). Summers et al. (1997) defined their Models B and C in the sixth paragraph of their text. Jucks et al. (1998) defined their Model D in the figure caption of their Figure 2; Model D1 is a code name we assign to the alternative model described in the eighth paragraph of Jucks et al.'s (1998) "Discussion" section. Conway et al. (2000) defined their Model B in the figure caption of their Plate 2. For this work, Experiments I–IV are defined in section 3.3. For better visualization, negative values are quoted in parentheses.
[a]Based on JPL1994 Evaluation (DeMore et al., 1994). [b]Based on JPL1994 Evaluation. Model H_2O, O_3, and temperature were constrained. [c]Based on JPL1997 Evaluation (DeMore et al., 1997). Model H_2O, O_3 (below 47 km), N_2O, NO_y, CH_4, Cl_y, and temperature were constrained. [d]Based on JPL2002 Evaluation. Model H_2O, O_3, N_2O, CO, and temperature were constrained. [e]Based on JPL2011 Evaluation. Model H_2O and temperature were constrained. [f]Based on JPL2011 Evaluation. [g]$k_{H2O+O(1D)}$ in JPL2006 (Sander et al., 2006) and JPL2011 (Sander et al., 2011) Evaluations are 5% larger than JPL2002 Evaluation (Sander et al., 2002) at 170 K. [h]k_{OH+O} in JPL2011 is 16% larger than JPL2006 and JPL2002 at 170 K. [i]k_{H+O2+M} in JPL2006 and JPL2011 are 36% less than JPL2002 at 170 K. [j]O_2 photo absorption cross section at Lyman α, resolved by a spectral resolution 0.1 nm at 121.57 nm in the 1-D photochemical model. [k]This experiment was not labeled in Jucks et al. (1998) but was discussed in the eighth paragraph of their Discussions section. [l]Conway et al. (2000) presented two other models: Models C and D, which were equivalent to Summers et al.'s (1997) Model C and Jucks et al.'s (1998) Model D, respectively. [m]Equivalent to 4% increase of JPL2011. [n]No observational constraints have been applied. [o]Model H_2O and O_3 between 30 and 72 km have been constrained by MLS observations. [p]Equivalent to JPL2002 and Wong and Davis (1974).

that accounts for both observational and model uncertainties to adjust the photochemical model parameters. Our approach is to set up an inverse problem, where parameters of the forward model (i.e., the photochemical model) would be estimated given the MLS observations. Below, we refer to the term "inversion" in a broad sense for the search of the model parameter values that would minimize a cost function defined in terms of observations and forward model outputs. When the observation is a set of satellite spectral measurements and the model parameter is the vertical profile of an atmospheric tracer (e.g., stratospheric ozone), for instance, then the search for the vertical profile is also known as "satellite retrieval" (Rodgers, 2000). In applied mathematics, an "inversion" described above is also known as an "optimization" of model parameters.

A number of studies (e.g., Canty et al., 2006; Siskind et al., 2013; Summers et al., 1997) attempted to relate the model-observation discrepancy in HO_x to the O_3 deficit problem, which is beyond the scope of this work. Our focus here is to apply an inversion method to atmospheric chemical modeling, emphasizing how to choose an appropriate set of reactions to be optimized and how to interpret the inversion results.

The rest of the paper is organized as follows. Section 2 presents the data and the photochemical model to be used in the inverse problem. Section 3 describes the inversion algorithm and presents the results. We will identify the need for reconsiderations of the laboratory data and recommended rate coefficients for $H + O_2 + M \rightarrow HO_2 + M$ (R12). Discussions and conclusions follow in section 4.

2. Data and Model

2.1. MLS Data

The MLS instrument aboard the Aura spacecraft was launched in 2004 (Waters et al., 2006) into a Sun-synchronous orbit, crossing the equator at around 1:45 A.M./P.M. We shall use the MLS version 4.2 daytime OH data (zonally and tropically averaged over 25°S and 25°N) during June 2005. This period was the first summer, having the strongest solar activity after launch; this choice of the period ensures that the measured HO_x concentrations have the highest signal-to-noise ratio. For the purpose of inversion, we interpolate the MLS profiles to the model levels that are uniformly separated by 2 km from ground to 120 km (see section 2.2). For quality control, the observed OH profile between 26 and 82 km (29 levels) and the observed HO_2 profiles between 38 and 82 km (23 levels) are used in the inversion.

We define a measurement vector by concatenating the OH and HO_2 profiles

$$\mathbf{y} = \begin{bmatrix} \mathrm{OH}(26\ \mathrm{km}) \\ \vdots \\ \mathrm{OH}(82\ \mathrm{km}) \\ \mathrm{HO_2}(38\ \mathrm{km}) \\ \vdots \\ \mathrm{HO_2}(82\ \mathrm{km}) \end{bmatrix}_{\mathrm{MLS}}. \tag{1}$$

The observational errors, $\sigma(z)$, are assumed to be 5% and 20% for OH and HO_2, respectively, based on the systematic errors of the MLS retrievals (Livesey et al., 2015). We define a 52×52 measurement error covariance matrix by concatenating the observational errors on the diagonal:

$$\mathbf{S}_e = \begin{bmatrix} \sigma^2_{\mathrm{OH}(26\ \mathrm{km})} & & & & & 0 \\ & \ddots & & & & \\ & & \sigma^2_{\mathrm{OH}(82\ \mathrm{km})} & & & \\ & & & \sigma^2_{\mathrm{HO_2}(38\ \mathrm{km})} & & \\ & & & & \ddots & \\ 0 & & & & & \sigma^2_{\mathrm{HO_2}(82\ \mathrm{km})} \end{bmatrix}. \tag{2}$$

The off-diagonal elements are all zeros.

2.2. The Photochemical Model

The Caltech/JPL 1-D photochemical model is used to simulate the stratospheric/mesospheric OH and HO_2 (Allen et al., 1981). This model contains 66 levels from the surface to 130 km altitude. Vertical transport is parameterized using eddy diffusion. The model contains 34 photolytic reactions and 142 bi-/ter-molecular reactions, including important reactions for odd-oxygen, odd-nitrogen, and odd-hydrogen species. The full list of reactions and their Reaction IDs can be found in the supporting information. Monthly mean solar flux data are from the Naval Research Laboratory (NRL) model (Lean, 2000).

For a bimolecular reaction, the rate coefficient is given by

$$k = A \exp\left(-\frac{E_a}{RT}\right) \tag{3}$$

where A is the reaction rate coefficient (also known as the collision frequency factor), E_a is the energy barrier for the reaction, T is the ambient temperature, and $R = 8.31\ \mathrm{J\ K^{-1}\ mol^{-1}}$ is the gas constant. The values of A and E_a/R are taken from sections 1 and 2 of 2011 JPL Data Evaluation (Sander et al., 2011). For a termolecular reaction, k is expressed in terms of high- and low-pressure limiting values and their temperature dependences.

Assuming that the HO_x lifetime depends mostly on the concentration of key species such as O, O_2, and O_3 and is much shorter than the transport time scales, we simply fix the profiles of air density, nitrogen gas, oxygen gas, and temperature with those in the U.S. Standard Tropical Atmosphere throughout the model calculations regardless of the time of the day. We note that a self-consistent diurnal cycle calculation is important for species such as HO_x and O_3, which have significantly different lifetimes at different altitudes. Thus, in the first set of experiments (Experiments I–III), all minor species, including H_2O, O_3, H_2, OH, and HO_2, are unconstrained and their diurnal cycles are calculated. This approach is different from previous studies,

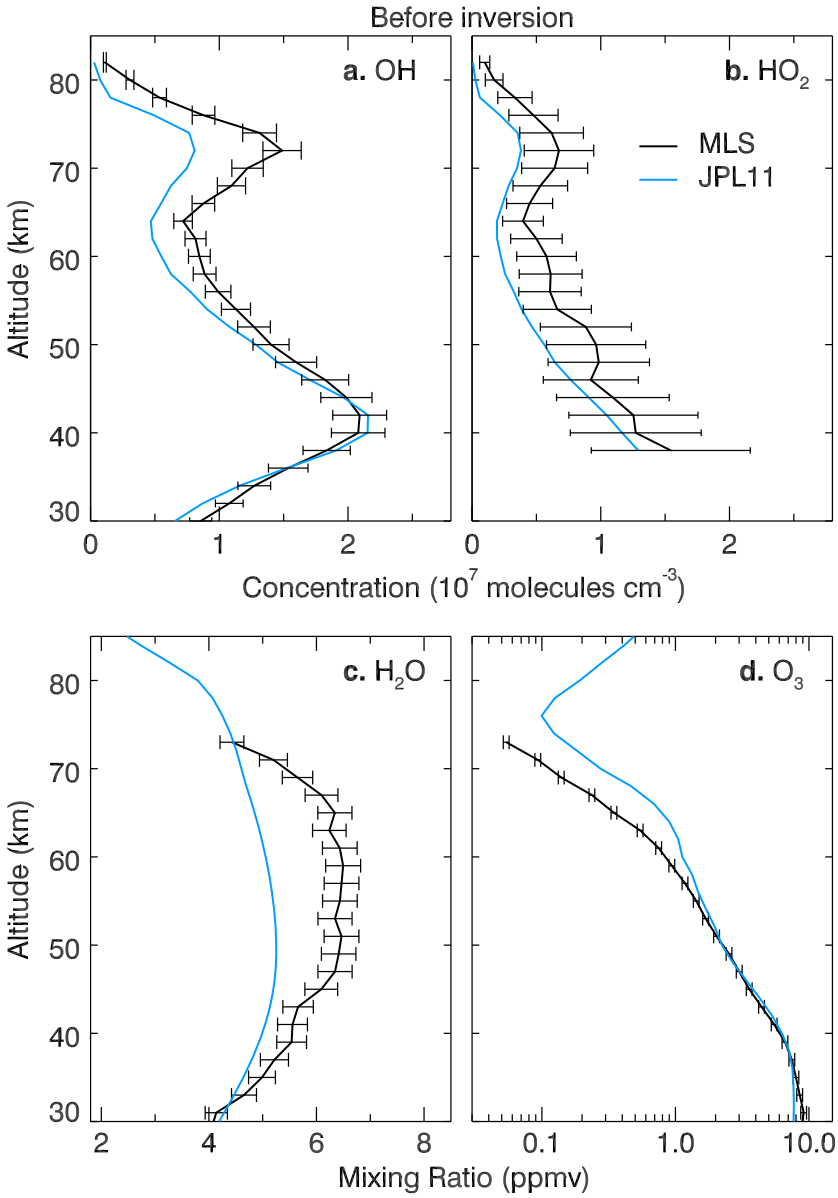

Figure 1. Comparison between MLS 1:45 P.M. measurements (black) and 1-D model simulations (blue): (a) OH, (b) HO_2, (c) H_2O, and (d) O_3. MLS daytime measurements are averaged between 25°S and 25°N from 1 to 30 June 2005. The blue curves show the model results using the kinetic rate coefficients from the 2011 JPL Data Evaluation. For visualization, H_2O and O_3 mixing ratios are shown. For OH, H_2O, and O_3, we assume a systematic error of 5%. For HO_2, we assume a systematic error of 20%.

where various observed concentrations of H_2O and O_3, as well as minor species such as N_2O, NO_y, CH_4, and Cl_y, have been used to constrain the model calculations; see the footnotes in Table 1. A problem with our unconstrained calculation is that the simulated H_2O and O_3 profiles may be different from the observed ones, as shown in Figures 1c and 1d: the simulated H_2O mixing ratio is at least 20% less than the MLS H_2O over altitudes between 44 and 72 km, and the simulated O_3 mixing ratio is about 50% less than the MLS O_3 between 60 km and 72 km. To illustrate the effect of the observational constraints on the retrieved kinetic rates, we perform another experiment (Experiment IV), where the model H_2O and O_3 profiles between 30 km and 72 km are fixed with the MLS observation. Experiment IV is similar to the model simulation conducted by Millán et al. (2015).

The model is run until the OH and HO_2 profiles at 1:45 P.M. over the equator (0°N) become steady. Analogous to **y**, we define a model vector $\mathbf{y}_m$ that concatenates the simulated OH and HO_2 profiles:

$$\mathbf{y}_m = \begin{bmatrix} \mathrm{OH}(26\ \mathrm{km}) \\ \vdots \\ \mathrm{OH}(82\ \mathrm{km}) \\ \mathrm{HO_2}(38\ \mathrm{km}) \\ \vdots \\ \mathrm{HO_2}(82\ \mathrm{km}) \end{bmatrix}_{\mathrm{model}}. \tag{4}$$

2.3. Discrepancies Between Observed and Model OH and HO_2 Profiles

Figures 1a and 1b show the monthly mean zonal-mean MLS OH and HO_2 profiles averaged between 25°S and 25°N in June 2005 (black lines). The use of a monthly mean minimizes the effects of the 27 day solar cycle. The simulated OH and HO_2 profiles (blue lines) are compared with the MLS observations. Three differences are noted. (1) The simulated mesospheric OH and HO_2 at 72 km are significantly underestimated, both only half of the observed, consistent with Millán et al.'s (2015) conclusion; (2) the simulated stratospheric OH at 40 km is slightly more than observed but they agree within the measurement uncertainty, consistent with Canty et al.'s (2006) conclusion; and (3) the simulated stratospheric HO_2 at 40 km is less than observed but they agree within the measurement uncertainty. Figures 1c and 1d show the simulated H_2O and O_3 profiles, as discussed in section 2.2. We assume a systematic error of 5% for both MLS H_2O and O_3 measurements.

Below we will test whether the model-observation differences can be improved by adjusting some of the reaction rates in the inversion.

3. Inversion of Kinetic Rates

Benefited from the simplicity of the HO_x photochemistry, we limit our parametric study to reaction rate constants and assume that the uncertainties of atmospheric transport and ambient temperature are not dominant. As in previous works (e.g., Canty et al., 2006; Pickett et al., 2008; Siskind et al., 2013), we will adjust the reaction rates A of some "important" chemical reactions for HO_x to minimize the discrepancy between the simulated and observed OH and HO_2 profiles. The same adjustment of A is applied to all model levels. E_a/R and T remain unchanged during the inversion.

3.1. Jacobians and Selection of Reactions

When the kinetic rate of a reaction is adjusted, the vertical profiles of OH and HO_2 will change. The partial derivative of the vertical profiles with respect to the kinetic rate defines the sensitivity of OH or HO_2 to the associated reaction. The collection of the partial derivatives form a 52×176 Jacobian matrix:

$$\mathbf{K} = \left[\frac{\partial \mathbf{y}_m}{\partial A_1} \cdots \frac{\partial \mathbf{y}_m}{\partial A_{176}}\right] = \begin{bmatrix} \frac{\partial [\mathrm{OH}]_{26\ \mathrm{km}}}{\partial A_1} & \cdots & \frac{\partial [\mathrm{OH}]_{26\ \mathrm{km}}}{\partial A_{176}} \\ \vdots & & \vdots \\ \frac{\partial [\mathrm{OH}]_{82\ \mathrm{km}}}{\partial A_1} & \cdots & \frac{\partial [\mathrm{OH}]_{82\ \mathrm{km}}}{\partial A_{176}} \\ \frac{\partial [\mathrm{HO_2}]_{38\ \mathrm{km}}}{\partial A_1} & \cdots & \frac{\partial [\mathrm{HO_2}]_{38\ \mathrm{km}}}{\partial A_{176}} \\ \vdots & & \vdots \\ \frac{\partial [\mathrm{HO_2}]_{82\ \mathrm{km}}}{\partial A_1} & \cdots & \frac{\partial [\mathrm{HO_2}]_{82\ \mathrm{km}}}{\partial A_{176}} \end{bmatrix}, \tag{5}$$

where A_i is the rate coefficient of the ith reaction. As an example, $\frac{\partial [\mathrm{OH}]_{26\ \mathrm{km}}}{\partial A_{\mathrm{OH+HO_2}}}$ is defined as the percent change in the OH concentrations at 26 km due to a 100% increase in the kinetic rate for the sink process $OH + HO_2 \rightarrow H_2O + O_2$.

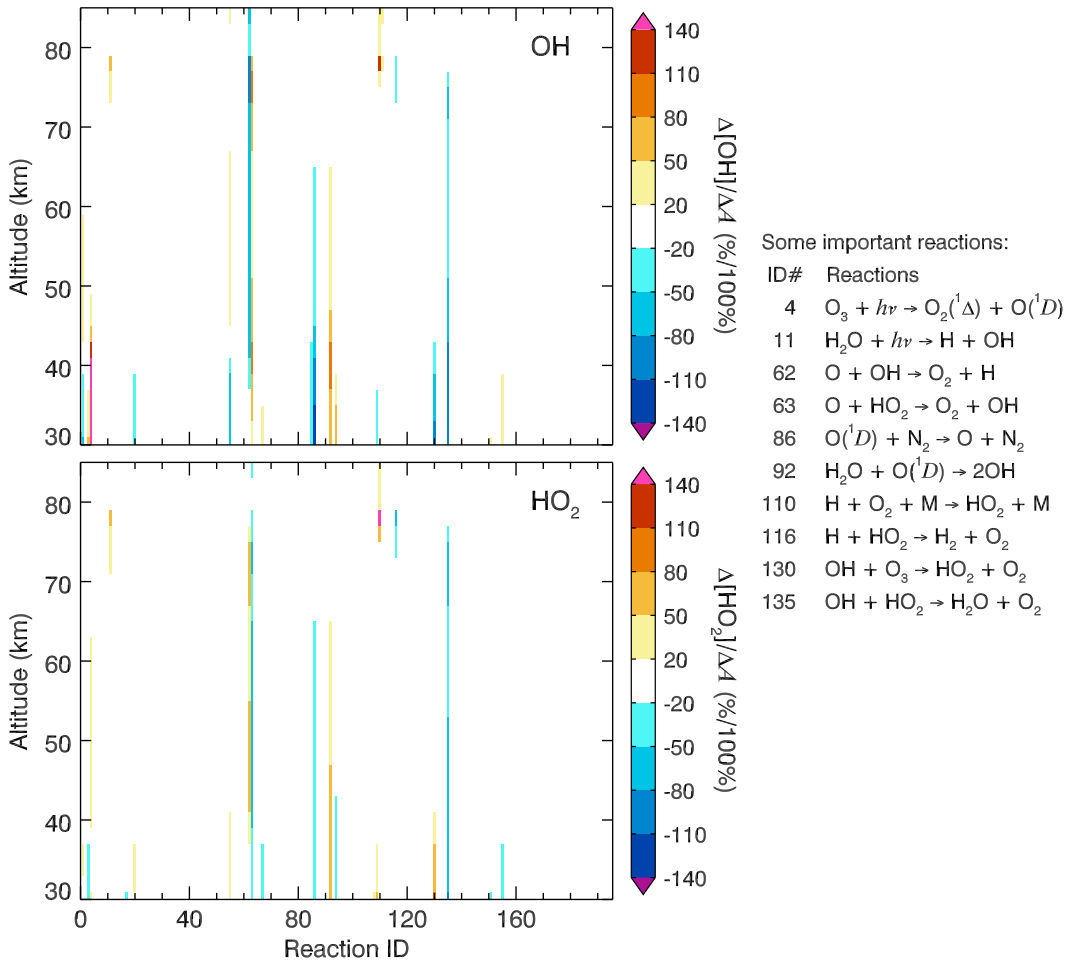

Figure 2. The Jacobian of (top) OH and (bottom) HO_2 with respect to reaction rate constants. The Jacobian at a particular altitude is defined as the percent change in OH or HO_2 concentration per 100% change in the reaction rate coefficient. Reaction ID are listed in the supporting information.

The selection of reactions to be adjusted is based on **K**, shown as partial derivatives of OH and HO_2 in Figure 2. The following HO_x reactions are found to have significant partial derivatives at 40 km and 72 km where the stratospheric and mesospheric peaks of HO_x are located: $O + HO_2 \rightarrow O_2 + OH$, $OH + HO_2 \rightarrow H_2O + O_2$, $O + OH \rightarrow O_2 + H$, $H + O_2 + M \rightarrow HO_2 + M$, and $H + HO_2 \rightarrow H_2 + O_2$. These reactions may be candidates for the inversion. However, only some of them can be used. This is not only because of the limited number of degrees of freedom to be derived in section 3.2 but also because of "inherent degeneracy" under the context of inversion: the partial derivatives of some reactions may have similar vertical structures, e.g., having peaks at similar altitudes. A simple example of inherent degeneracy is the two reactions among odd-oxygen species and oxygen gas that have almost symmetrically opposite effects on both OH and HO_2: $O + O_2 + M \rightarrow O_3 + M$, $O + O_3 \rightarrow 2O_2$. The adjustments of values of *A* for these two reactions would not be unique, albeit the same changes in OH and HO_2 may be obtained. Therefore, in the selection process, we need to avoid selecting reactions that have structurally similar Jacobians. As the reactions $O + OH \rightarrow O_2 + H$ and $O + HO_2 \rightarrow O_2 + OH$ have similar Jacobians (Figures 3a and 3b), only one of them should be used in the inversion. For the same reason, since the reaction $H + HO_2 \rightarrow H_2 + O_2$ has a Jacobian (Figure 3c) that is similar to that of $H + O_2 + M \rightarrow HO_2 + M$ (not shown), we will not include the former reaction in the inversion.

3.2. Bayesian Optimal Estimation

We adopt the Bayesian optimal estimation to retrieve the reaction rate coefficients (Rodgers, 2000). The recommended reaction rate coefficients in the 2011 JPL Data Evaluation will be used as the a priori values

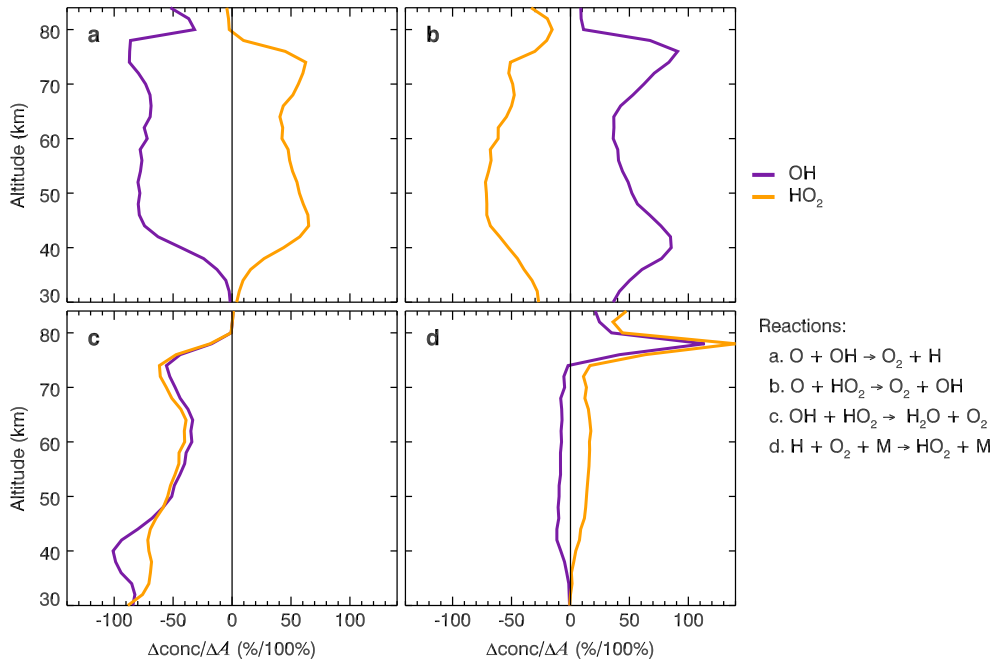

Figure 3. The Jacobians of OH (purple) and HO_2 (orange) with respect to the labeled kinetic rate coefficients.

for A. The a priori uncertainties of A are given by $[f(T)-1]\times 100\%$, where $f(T)$ is an uncertainty scaling function defined in sections 1.2 and 2.6 of Sander et al. (2011), and $f(T)$ generally lies within 10–30% (but it can be as large as 200% for some reactions). We define a diagonal a priori error covariance matrix as

$$\mathbf{S}_a = \begin{bmatrix} (f_1-1)^2 & & 0 \\ & \ddots & \\ 0 & & (f_{176}-1)^2 \end{bmatrix}. \tag{6}$$

One cannot choose an arbitrary number of reactions for inversion because the MLS observations may not have sufficient degrees of freedom to constrain a large number of model parameters. The degree of freedom of the MLS measurements is defined as (Rodgers, 2000)

$$d_s = \mathrm{Tr}\left(\mathbf{K}\mathbf{S}_a\mathbf{K}^T\left[\mathbf{K}\mathbf{S}_a\mathbf{K}^T+\mathbf{S}_e\right]^{-1}\right), \tag{7}$$

where Tr(**M**) is the trace, i.e., the sum of the diagonal elements, of a square matrix **M**. d_s measures how many model parameter can be adjusted independently when fitting the MLS OH and HO_2 profiles. As a rough estimate, we assume $(f-1)\approx 20\%$ for all reactions and obtain $d_s = 6.84$. Therefore, no more than seven reactions can be adjusted in the optimal estimation.

After the selection of reactions, the reaction rates of the selected reactions are to be adjusted such that the following Bayesian cost function is minimized

$$\chi^2 = [\mathbf{y}-\mathbf{y}_m(\mathbf{x})]^T\mathbf{S}_e^{-1}[\mathbf{y}-\mathbf{y}_m(\mathbf{x})] + (\mathbf{x}-\mathbf{x}_a)^T\widetilde{\mathbf{S}}_a^{-1}(\mathbf{x}-\mathbf{x}_a), \tag{8}$$

where $\mathbf{x}$ is a state vector containing the reaction rate coefficients to be adjusted, $\mathbf{x}_a$ contains the a priori reaction rate coefficients corresponding to those in $\mathbf{x}$, $\widetilde{\mathbf{S}}_a$ is the reduced a priori covariance

Table 2
The A Priori Uncertainties (1σ) of the Reaction Coefficients to be Adjusted

Reaction	Prior uncertainty
$O + OH \rightarrow O_2 + H$	15%
$OH + HO_2 \rightarrow H_2O + O_2$	15%
$H + O_2 + M \rightarrow HO_2 + M$	30%
O_2 photo absorption at Lyman α	30%

Note. The uncertainties of $O + OH \rightarrow O_2 + H$, $OH + HO_2 \rightarrow H_2O + O_2$, and $H + O_2 + M \rightarrow HO_2 + M$ are based on the JPL 2011 Evaluation (Sander et al., 2011). The uncertainty for the O_2 photo absorption cross section at Lyman α is heuristically derived; see text.

matrix with the selected reactions only, and $\mathbf{y}_m(\mathbf{x})$ is $\mathbf{y}_m$ given $\mathbf{x}$. For example, if $O + OH \rightarrow O_2 + H$ and $O + HO_2 \rightarrow O_2 + OH$ are to be adjusted, then $\mathbf{x} = [A_{O+OH}, A_{OH+HO_2}]$, and $\widetilde{\mathbf{S}}_a = \begin{bmatrix} (f_{O+OH}-1)^2 & 0 \\ 0 & (f_{OH+HO_2}-1)^2 \end{bmatrix}$. A reduced Jacobian $\widetilde{\mathbf{K}}$ can be similarly defined: $\widetilde{\mathbf{K}} = \left[\frac{\partial \mathbf{y}_m(\mathbf{x})}{\partial A_{O+OH}} \quad \frac{\partial \mathbf{y}_m(\mathbf{x})}{\partial A_{OH+HO_2}} \right]$.

The minimization of χ^2 is obtained by the Levenberg-Marquardt algorithm with the modification suggested by Nielsen (1999).

3.3. Inversion Results

To demonstrate the importance of the selected reactions, we present three experiments, where we progressively include more reactions to show the improvements due to the included reactions. In the first experiment, we reexamine the reactions chosen by Canty et al. (2006), which are $OH + HO_2 \rightarrow H_2O + O_2$ and $O + OH \rightarrow O_2 + H$. The second experiment add one more reaction chosen by Siskind et al. (2013), which is $H + O_2 + M \rightarrow HO_2 + M$. In the third experiment, we will propose our solution to the model-observation discrepancy, which is the O_2 absorption cross section at Lyman α (121.57 nm). The fourth experiment is to test the robustness of our inversion results with a constrained model where the observed H_2O and O_3 profiles are imposed.

3.3.1. Experiment I—Canty Et Al.'s Reactions

Canty et al. (2006) adjusted the reaction rates of $OH + HO_2 \rightarrow H_2O + O_2$ (loss of HO_x) and $O + OH \rightarrow O_2 + H$ (HO_x interconversion); see Table 1. Both reactions have very similar Jacobians in the mesosphere and the stratosphere, which implies degeneracy. The a priori uncertainties are $(f_{O+OH} - 1) = (f_{OH+HO_2} - 1) = 15\%$ (Table 2). The cost function χ^2 is minimized if A_{OH+HO_2} is increased by 29% and A_{O+OH} is reduced by 53%, which well exceed their a priori uncertainties. These results are also different from Canty et al.'s (2006) results, where A_{OH+HO_2} is increased by 20% and A_{O+OH} is increased by 4% (relative to the 2011 JPL Data Evaluation).

The resultant OH and HO_2 profiles are shown in Figure 4. There is a significant increase (~50%) in the mesospheric OH, caused by the much slower OH destruction through $O + OH \rightarrow O_2 + H$ after inversion. However, the same reaction leads to an increase in the stratospheric OH. To compensate this stratospheric increase, the HO_x sink reaction $OH + HO_2 \rightarrow H_2O + O_2$ has to be 29% faster. But the resultant OH concentration at 40 km still exceeds the observed value. More seriously, the faster HO_x sink reaction results in a significant decrease of HO_2 at all altitudes, leading to a larger discrepancy between the model and the observation.

As a result of the inversion, there is a slight increase in the simulated H_2O that makes the simulated H_2O profile be more consistent with the observed H_2O profile. However, there is also an increase in the simulated O_3 that moves the simulated O_3 profile further away from the observed O_3 profile.

As a sensitivity test on the choice of reactions having similar Jacobians, we perform another experiment where we replace $O + OH \rightarrow O_2 + H$ by $O + HO_2 \rightarrow HO + O_2$ (not shown), which was another HO_x conversion cycle chosen by Jucks et al. (1998). The improvement in the mesospheric OH is worse because of a more stringent a priori uncertainty $(f_{O+HO_2} - 1) = 5\%$.

3.3.2. Experiment II—Addition of $H + O_2 + M \rightarrow HO_2 + M$

Siskind et al. (2013) adopted a rate coefficient for $H + O_2 + M \rightarrow HO_2 + M$ from Wong and Davis (1974), which was 36% faster than the recommended value in 2011 JPL Data Evaluation in order to increase the mesospheric OH and HO_2. We add this reaction to Canty et al.'s (2006) selection and perform the inversion, with $(f_{H+O_2+M} - 1) = 30\%$ (Table 2). The Jacobian of this reaction has a singular peak at 76 km (Figure 3d). As a result, the cost function χ^2 is minimized if A_{OH+HO_2} is reduced by 13%, A_{O+OH} is reduced by 31%, and A_{H+O_2+M} is increased by 131%. In this experiment, the change in A_{OH+HO_2} is of opposite sign to that in Experiment I but is within the laboratory uncertainty. In contrast, the a posteriori A_{O+OH} and A_{H+O_2+M} well exceed the a priori uncertainties. A_{H+O_2+M} is more than double the a priori value and is much more than the rate coefficient suggested by Siskind et al. (2013) and Wong and Davis (1974) (Table 1).

The increased kinetic rate coefficient A_{H+O_2+M} help improve the simulated OH and HO_2 above 76 km significantly (Figure 5) by producing more HO_2 and hence OH (through $O + HO_2 \rightarrow O_2 + OH$) above

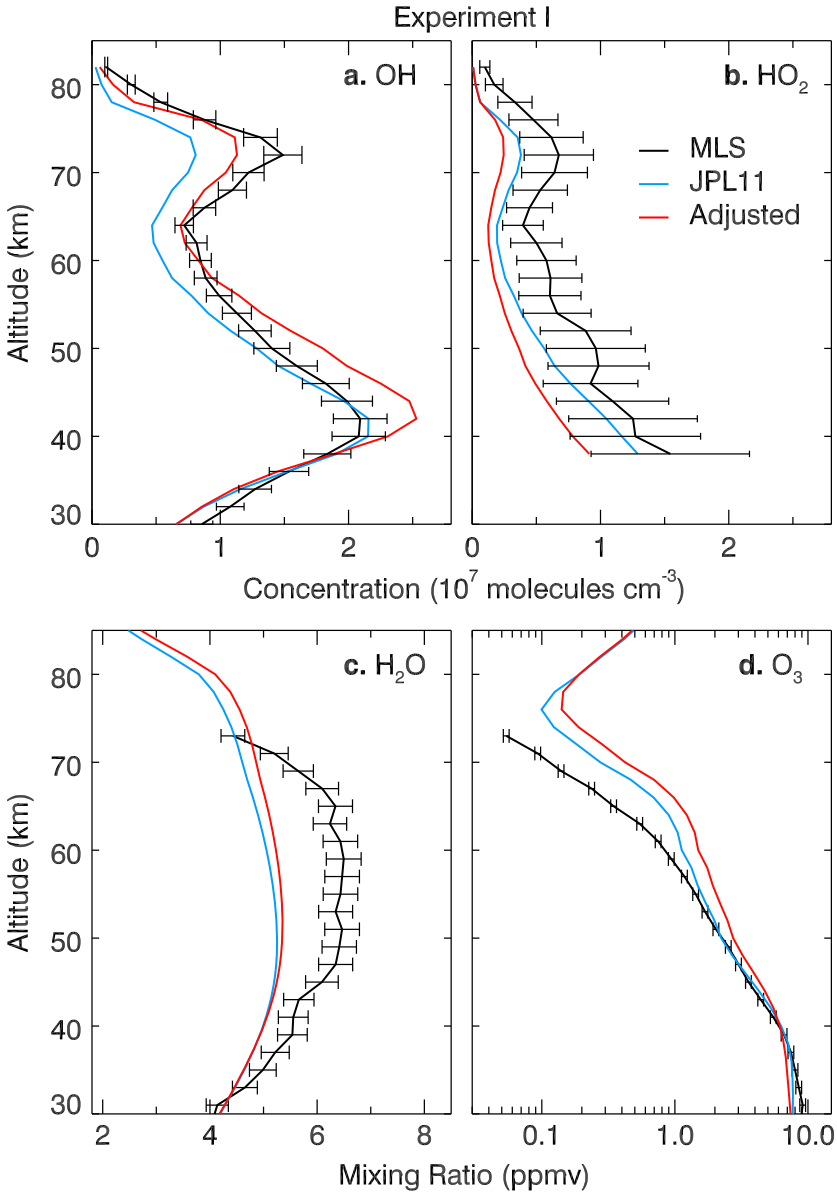

Figure 4. Same as Figure 1 except the simulated OH, HO_2, H_2O, and O_3 profiles after adjusting reaction rate coefficients for $O + OH \rightarrow O_2 + H$ and $OH + HO_2 \rightarrow H_2O + O_2$ are also shown.

72 km. As a result, the reduction in the kinetic rate of $O + OH \rightarrow O_2 + H$ is less than that in Experiment I. The modified OH profile below 72 km is the same as that in Experiment I, but the HO_2 profile below 72 km is the same as that before the inversion.

The doubling of A_{H+O_2+M} seems to contradict with the laboratory measurements. We shall discuss the important implications of the adjusted A_{H+O_2+M} in section 4.

As in Experiment I, there is an increase in the simulated H_2O, which makes the resultant H_2O profile more consistent with the observation. The simulated O_3, in contrast, remains the same as in a priori profile, except above 76 km, where the adjustment of A_{H+O_2+M} reduces the mesospheric O_3.

3.3.3. Experiment III—O_2 Absorption Cross Section

The improvements of mesospheric OH and HO_2 in Experiments I and II are not satisfactory. Parameters that have been adjusted in previous work are predominantly kinetic rates. An exception is Siskind et al. (2013),

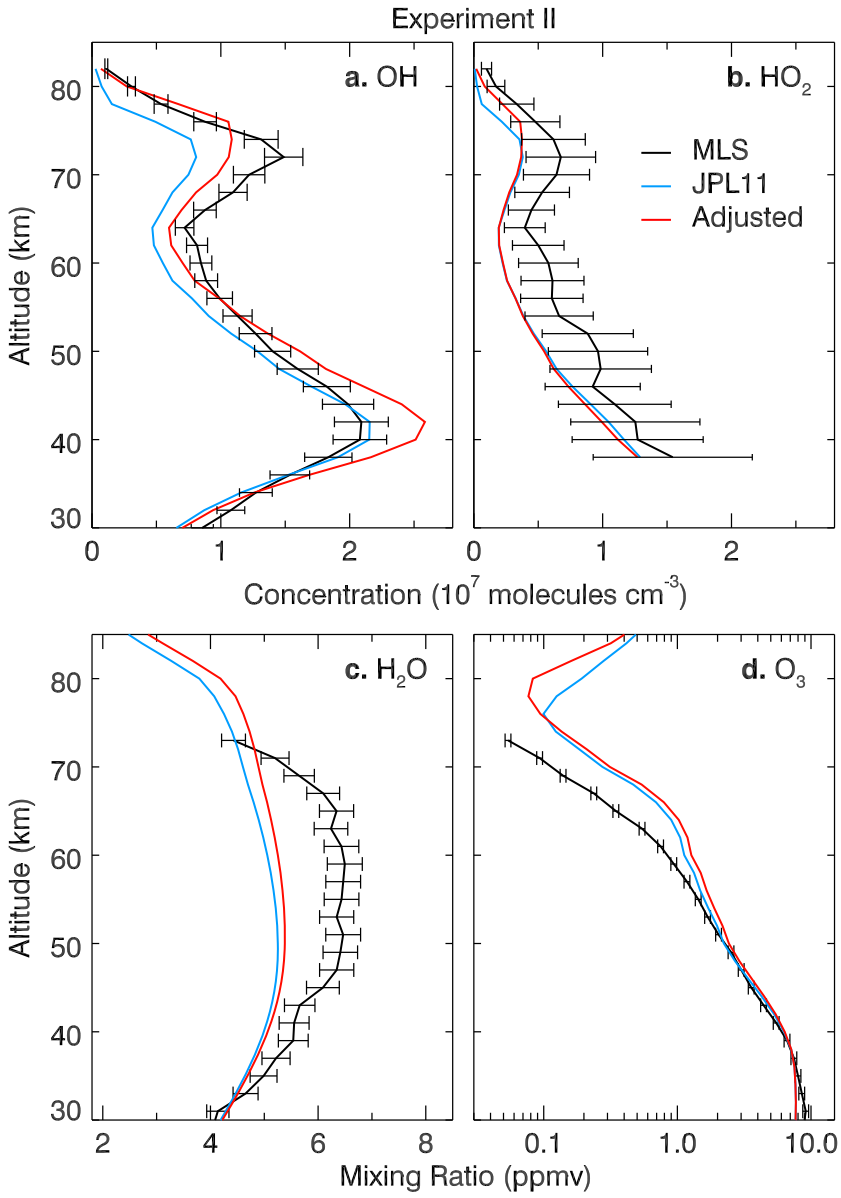

Figure 5. Same as Figure 4 except with adjustments of the reaction rate coefficients for O + OH → O_2 + H, OH + HO_2 → H_2O + O_2, and H + O_2 + M → HO_2 + M.

who also considered the uncertainty in the top-of-atmosphere (TOA) Lyman α intensity, which primarily affects the H_2O photo dissociation through H_2O + $h\nu$ (121.57 nm) → H + OH between 65 km and 80 km. We perform an experiment adjusting the Lyman α intensity alone (not shown). The Lyman α intensity would need to be increased by an unrealistic 300% to produce a mesospheric OH concentration comparable to the MLS observation. Therefore, we seek another solution.

The optical depth at Lyman α reaches unity above 80 km due to O_2 absorption (see Figure 3.7 of Liou, 2002). The weaker the O_2 absorption at Lyman α is, the deeper the solar Lyman α can penetrate into the mesosphere, and the stronger response of the H_2O photo dissociation to the 11 year solar variability will be. Thus, besides adjusting the TOA Lyman α intensity as Siskind et al. (2013) did, the H_2O photo dissociation rate at 75 km can also be indirectly adjusted by modifying the O_2 absorption cross section at Lyman α. To demonstrate this effect, we show in Figure 6a the partial derivatives of the OH and HO_2 profiles with respect to the

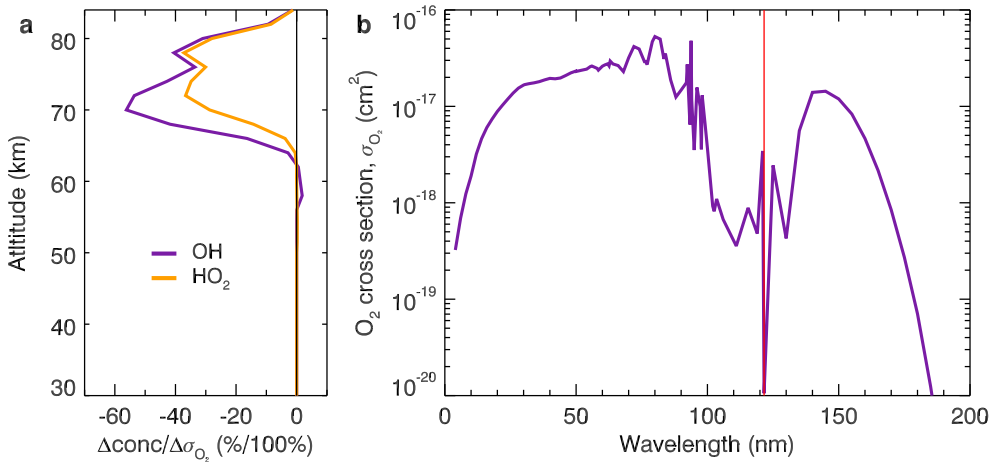

Figure 6. (a) Jacobian of OH (black) and HO_2 (red) with respect to the O_2 photo absorption cross section at Lyman α, in unit of percent change in concentration per 100% change in the O_2 photo absorption cross section (σ_{O_2}). (b) O_2 photo absorption cross section as a function of wavelength, Lyman α (121.57 nm) wavelength is marked with a vertical red line.

O_2 absorption cross section at the Lyman α line, which reveal a broad peak coinciding with the OH and HO_2 mesospheric peak.

Indeed, the O_2 photo absorption cross section at Lyman α is a singular dip, which is 2 orders of magnitude smaller than the continuum at the neighboring wavelengths (Figure 6b) (Liang et al., 2007). This "singular window" in the O_2 absorption cross section allows more Lyman α intensity to penetrate into the lower atmosphere than other FUV flux (Nicolet & Peetermans, 1980). We argue that the singular dip of the O_2 photo absorption cross section at Lyman α is very difficult to be measured accurately in the laboratory and may thus be subject to large uncertainty. We thus include the O_2 absorption cross section at Lyman α as another adjustable model parameter. From 121.52 nm to 121.62 nm, the O_2 cross section monotonically decreases from 1.74×10^{-20} cm^2 to 0.52×10^{-20} cm^2 (Lewis et al., 1983), while in the 1-D photochemical model, we adopt an average value of 1.06×10^{-20} cm^2 at 121.57 nm over this wavelength range.

The temperature dependence of the O_2 absorption cross section may be a source of uncertainty. Lewis et al. (1983) measured the O_2 absorption cross section near Lyman α at a very high resolution (0.01 nm). They showed that the implied column O_2 dissociation rate generally was within 10–15% of previously reported values if the temperature dependence was ignored. Another source of the uncertainty of the O_2 absorption cross section is due to the relatively coarse spectral resolution (0.1 nm) at Lyman α in our 1-D photochemical model, which is not enough to accurately represent the dramatic change up to several orders of magnitude (Figure 6b) (Ogawa, 1968). With the above considerations, we heuristically assume a conservative estimate of an uncertainty of 30% for the O_2 absorption cross section at Lyman α (Table 2).

As a result, the cost function χ^2 is minimized if the O_2 absorption cross section at the Lyman α line is reduced by 54%, A_{OH+HO_2} and A_{O+OH} are both reduced by 15%, and A_{H+O_2+M} is increased by 134%. Note that with the addition of O_2 absorption cross section, the adjustment in A_{O+OH} is now within the measurement uncertainty.

Figure 7 shows the OH and HO_2 profiles after the inversion. The mesospheric OH concentration between 60 and 74 km has significantly increased compared to the previous two experiments, which greatly improves the agreement between the simulated and observed OH mesospheric concentrations. Because the reduction of A_{O+OH} is less than that in Experiment II, the simulated stratospheric OH agrees better with observation. Meanwhile, the simulated mesospheric HO_2 also agrees better with observations.

As in Experiments I and II, there is an increase in the simulated H_2O. The reduction of the O_2 absorption cross section reduces the mesospheric O_3 between 62 km and 76 km and the simulated O_3 profile agrees better with the observed O_3 over this range.

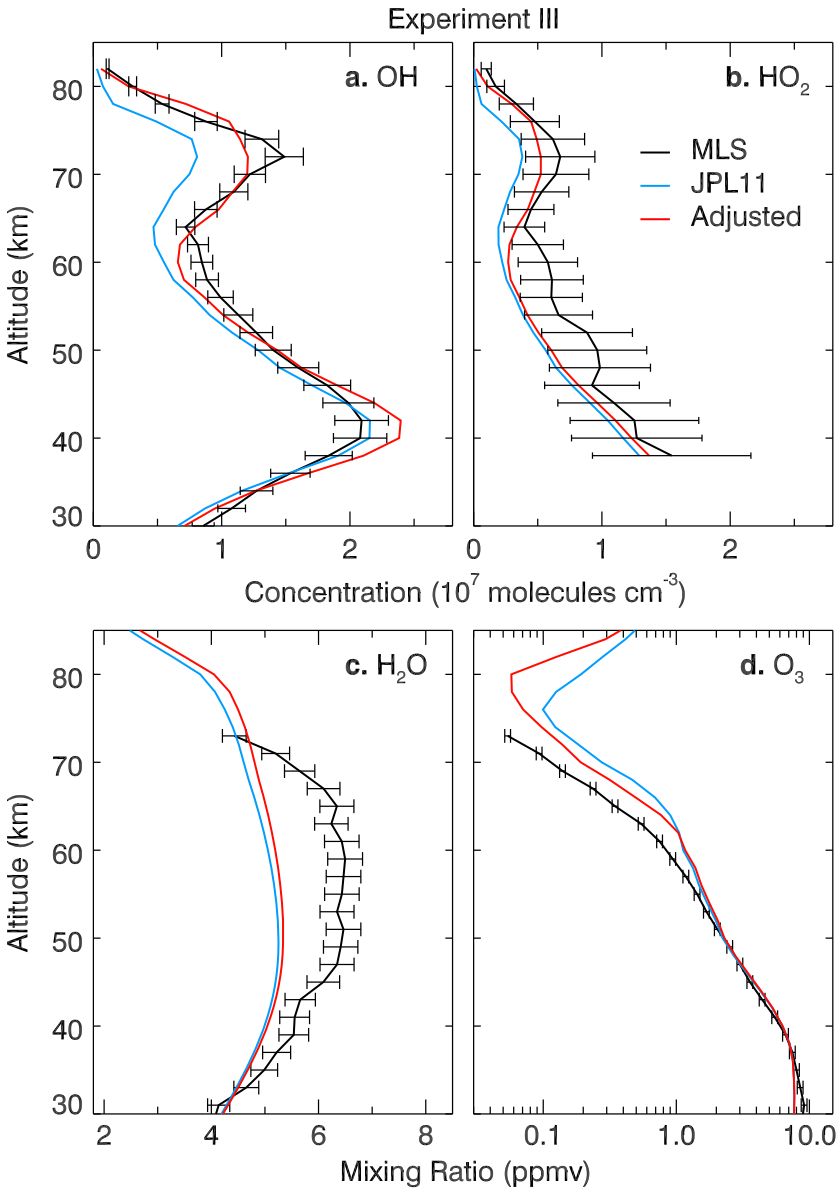

Figure 7. Same as Figure 4 except with adjustments of the reaction rate coefficients for O + OH → O_2 + H, OH + HO_2 → H_2O + O_2, H + O_2 + M → HO_2 + M, and the O_2 absorption cross section at Lyman α.

Again, as a sensitivity test on the choice of reactions having similar Jacobians, we replace O + OH → O_2 + H by O + HO_2 → O_2 + OH and perform the inversion. In this case, the cost function χ^2 is minimized if the O_2 absorption cross section at the Lyman α line is reduced by 58%, A_{OH+HO_2} is reduced by 30%, A_{O+HO_2} is reduced by 10%, and A_{H+O_2+M} is increased by 173%. Thus, the inversion results are qualitatively consistent with those by adjusting A_{O+OH}.

3.3.4. Experiment IV—Simulations With H_2O and O_3 Constraints

A concern about Experiments I–III is that the modeled H_2O and O_3 profiles are not realistic comparing to the MLS observations (see the discussion in section 2.2 and see Figures 1, 4, 5, and 7). We thus recalculate the model OH and HO_2 profiles by constraining the H_2O and O_3 profiles using the zonally and tropically (25°N–25°S) averaged MLS daytime observation at 30–72 km. The a priori profiles are plotted in Figure 8 (green lines), which are compared with the a priori profiles obtained from the unconstrained simulation (blue lines). After imposing the MLS H_2O and O_3 constraints, the simulated OH and HO_2 concentrations are

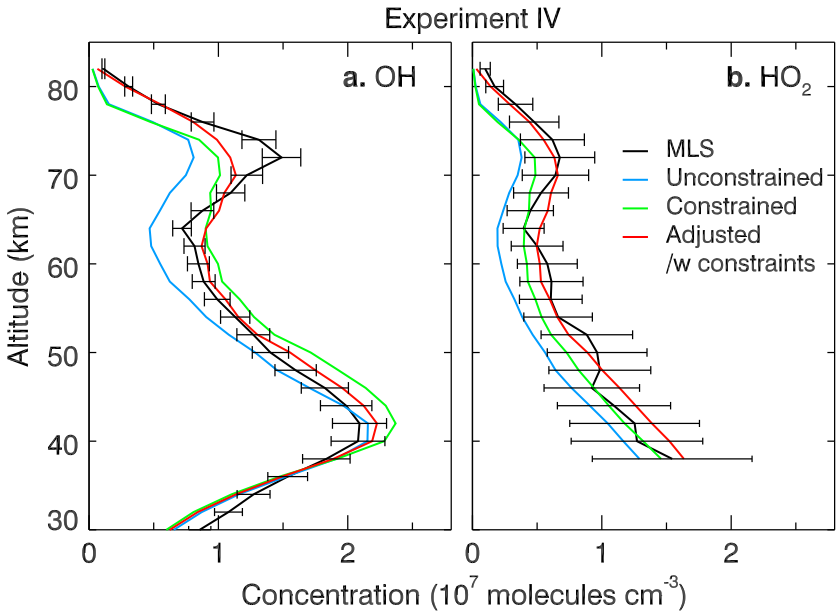

Figure 8. Comparison of the simulated OH and HO_2 profiles (i) before (blue) and after (green) constraining the model H_2O and O_3 profiles with the MLS observations and (ii) before (green) and after (red) the adjustments of the reaction rate coefficients for $O + OH \rightarrow O_2 + H$, $OH + HO_2 \rightarrow H_2O + O_2$, $H + O_2 + M \rightarrow HO_2 + M$, and the O_2 absorption cross section at Lyman α given the MLS H_2O and O_3 constraints.

generally higher between 40 and 72 km than the unconstrained simulation, primarily due to the increased source of OH from the photolysis of H_2O and the recombination of H_2O and $O(^1D)$. Moreover, while the simulated OH is higher than observed below 65 km, it is significantly lower than observed above 65 km, especially at the mesospheric peak, which is consistent with the result of Millán et al. (2015).

With the MLS H_2O and O_3 constraints, the cost function χ^2 is minimized if the O_2 absorption cross section at the Lyman α line is reduced by 34%, A_{OH+HO_2} is reduced by 10%, A_{O+OH} is increased by 12%, and A_{H+O_2+M} is increased by 310%. The a posteriori OH and HO_2 profiles are shown in Figure 8 (red lines). Three comments are in order.

1. The inferred change in the O_2 absorption cross section at Lyman α is qualitatively consistent with the result obtained from Experiment III. Therefore, the application of the MLS H_2O and O_3 constraints still implies the need for a reexamination of the O_2 absorption cross section at Lyman α.
2. A_{H+O_2+M} is still required to be much larger than the 2011 JPL Data Evaluation after imposing the MLS H_2O and O_3 constraints, again urging for a reexamination of the reaction $H + O_2 + M \rightarrow HO_2 + M$.
3. Although the change of A_{O+OH} becomes positive after imposing the MLS H_2O and O_3 constraints, the magnitude of the change is well within the experimental uncertainty. Thus, the inferred A_{O+OH} is consistent with the 2011 JPL Data Evaluation. The increase in A_{O+OH} leads to a decrease the stratospheric OH and an increase the stratospheric HO_2, so that the adjusted stratospheric OH and HO_2 agree better with the observations than in other experiments.

4. Summary and Discussion

We have proposed a systematic approach to estimate model parameters based on high-quality satellite observations. The Bayesian optimal estimation helps quantify model parameter uncertainties and provide guidance to laboratory measurements for key reactions. Such an inversion requires a large number of runs to estimate the model sensitivity with respect to each parameter. A computationally inexpensive 1-D photochemical model is particular useful for such a sensitivity calculation. Since HO_x chemistry in the mesosphere and stratosphere is simple and mainly controlled by several key reactions, the simplification in the transport in the 1-D model has little effect on our conclusions. In this model, all transports including vertical winds and

gravity wave mixing (Grygalashvyly et al., 2011) are parameterized using eddy diffusion. The most significant impact on the concentrations of HO_x species from this simplified scheme is the transport of H_2O. We tested the impact of perturbations in the vertical eddy diffusivity to the OH and HO_2 profiles. The largest impact is found in the lower stratospheric HO_x, due to the transport of tropospheric H_2O into the stratosphere at the tropopause which subsequently enhanced the production OH via $H_2O + O(^1D) \rightarrow 2OH$. However, the partial derivatives of the OH and HO_2 profiles with respect to changes in the eddy diffusivity is of the order of 10^5, which is one order of magnitude smaller than the Jacobians of important reaction rate coefficients.

We have used the kinetic rate coefficients in the 2011 JPL Data Evaluation, which may differ from previous evaluations; see the footnotes in Table 1. For example, the 2011 JPL Data Evaluation for the kinetic rate coefficient for $O + OH \rightarrow O_2 + H$ is 16% higher than that in the 2006 JPL Data Evaluation. Thus, the retrieved rate for $O + OH \rightarrow O_2 + H$ in this study would be 31% larger than that in JPL 2006 Evaluation. We have considered both the stratospheric and mesospheric peaks in the OH and HO_2 profiles, in contrast to some previous studies, such as Canty et al. (2006), who considered only stratospheric OH and HO_2. In addition to the reactions listed in Table 1, we also tried other combinations of reactions, including some NO_x reactions. Their fittings are worse than the result we have shown above.

The kinetic rate coefficient for $H + O_2 + M \rightarrow HO_2 + M$ is required to be more than double, while the measurement uncertainty recommended by the 2011 JPL Data Evaluation is only 30%. This is also much larger than the perturbations made to other model parameters. There are two possible explanations. The first one relies on the fact that this reaction is primarily contributing to the production of OH and HO_2 in the mesosphere (Figure 3b). At this altitude, pressure and temperature are extremely low. At 78 km where the Jacobian for this reaction is maximum, atmospheric pressure is only 0.022 hPa. Most of the measurements of this reaction are done at much higher pressure and temperature (Sander et al., 2011, and references therein) and may not be accurate at such a low pressure level, where the characteristic times of odd hydrogen and odd oxygen increase with altitude and the reactions at high altitude may not have reached the steady state. Thus, in view of the large increase in rate constant for the $H + O_2 + M$ inferred from our inversion, we have reexamined the kinetics database for this reaction, similar to that presented in Siskind et al. (2013). The NASA JPL Panel considered 11 laboratory studies of this reaction which used several different techniques over a wide range of pressure and temperature. A large majority of these studies focused on the temperature range relevant to combustion conditions, $298 \leq T \leq 1500$ K. Only two studies presented data relevant to the middle atmosphere. Both Kurylo (1972) and Wong and Davis (1974) used the flash photolysis-atomic resonance fluorescence technique to measure termolecular rate coefficients below room temperature using several different bath gases. For $M = N_2$ at 220 K, Kurylo (1972) obtained 8.35×10^{-32} cm^6 molecule^{-2} s^{-1}, while Wong and Davis (1974) obtained $(8.6 \pm 1.6) \times 10^{-32}$ cm^6 molecule^{-2} s^{-1}. The rate coefficient recommended by the NASA Panel for $M = N_2$ at 220 K is considerably smaller: 6.6×10^{-32} cm^6 molecule^{-2} s^{-1}. At 298 K, where there are several additional studies, the average of the k_{298} rate coefficients is also about 25–30% larger than the NASA recommendation. It should be noted that the NASA Panel accepted the recommendation contained in a theoretical paper by Sellevåg et al. (2008) which was aimed at obtaining a suitable fit between two-dimensional master equation calculations and the high-temperature kinetics database for the purposes of combustion studies. Inspection of Figure 4b in Sellevåg et al. (2008) which compares their master equation results with the lab data near room temperature for $M = N_2$ clearly shows that the theoretical results fall below all the experimental data in the termolecular pressure regime. Therefore, the Sellevåg et al. (2008) values, and implicitly the JPL Data Evaluation, is unsuitable for the pressure and temperature range of interest for the altitude regime considered in the present study. At 170 K in the lower mesospheric region, the rate coefficient calculated using Wong and Davis's (1974) exponential parameterization that is 52% larger than the 2011 JPL Data Evaluation and Siskind et al. (2013) showed that this rate coefficient produced a more realistic HO_x concentration in the lower mesosphere. Therefore, we suggest that a value for the $H + O_2 + N_2$ termolecular rate coefficient that is 25−50% larger than the JPL Data Evaluation is an appropriate choice. Yet this increase in the rate coefficient is still too small compared to our optimized value.

A second possible explanation, which might help enhance the effective rate constant of $H + O_2 \rightarrow HO_2$ under upper stratospheric conditions and which has not been considered in the literature thus far, is the radiative association (Vuitton et al., 2012). In the mesosphere, the limiting factor of the three body reaction is the total concentration of M due to the low pressure. At this level, the radiative association reaction ($A + B \rightarrow AB + h\nu$)

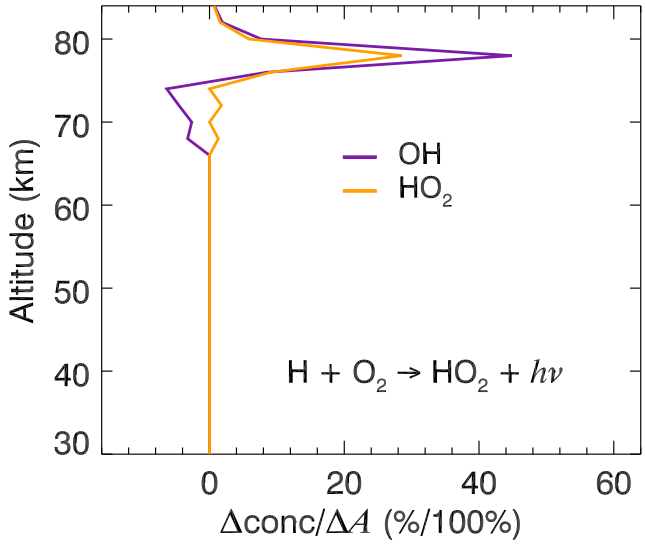

Figure 9. The Jacobians of OH (purple) and HO_2 (orange) with respect to the radiative association reaction $H + O_2 \rightarrow HO_2 + h\nu$ with a nominal reaction rate of 7×10^{-17} s^{-1} cm^3.

may have a similar or even higher order of magnitude reaction rate than the three-body reaction (A + B + M → AB + M). The reaction enthalpy, ΔH (298 K), for the radiative association is –49.2 kcal mol^{-1}, which is much larger than the energy required to populate the low-lying $\widetilde{A}$ electronic state of HO_2 at about 17,200 cm^{-1} provided that the required electronic curve crossing is sufficiently rapid. This would also require a favorable fluorescence lifetime for the $\widetilde{A}^2A' \rightarrow \widetilde{X}^2A''$ transition, which is reasonably strong in absorption. As an estimation, we calculate the reaction rate of $H + O_2 + M \rightarrow HO_2 + M$ at 78 km altitude using the low-pressure limit:

$$k = k_0^{300}\left(\frac{T}{300}\right)^{-n}[M] \quad (9)$$

where T = 191.6K, $[M] = 6.32\times10^{14}$ cm^{-3}. In the JPL 2011 evaluation, $k_0^{300} = 4.4 \times 10^{-32}$ s^{-1} cm^6, n = 1.3. To compensate for the 134% increase in the three body reaction rate for $H + O_2 + M \rightarrow HO_2 + M$, a radiative association reaction rate coefficient of ~7×10^{-17} s^{-1} cm^3 for $H + O_2 \rightarrow HO_2 + h\nu$ is needed. While the radiative association reaction rates have never been measured, Vuitton et al. (2012) calculated several radical-molecule reaction rates theoretically using transition state theory. They found that the contribution of the photo association reaction in a two-heavy-atom radical-radical reaction rate coefficient is in the order of 1.0×10^{-17} s^{-1} cm^3. As a test, this reaction is added to our 1-D model with a nominal reaction rate of 7×10^{-17} cm^3 s^{-1}. The partial derivatives of OH and HO_2 with respect to this new reaction exhibit the same sharp peaks in the mesosphere compared with those of $H + O_2 + M \rightarrow HO_2 + M$ (Figure 9). The values of their partial derivatives also have the same order of magnitude as those of $H + O_2 + M \rightarrow HO_2 + M$.

The method proposed in this study is not limited to studying HO_x chemistry. We choose to use the stratospheric and mesospheric HO_x mean profiles because they are very well measured by MLS and are mainly controlled by simple chemistry. The same method could be applied to solve other model-observation discrepancy problems.

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Earch and Space Science

Supporting Information for

Resolving Model-Observation Discrepancy in the Mesospheric and Stratospheric HO_x Chemistry

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Text S1

Complete photolytic and chemical reaction list with Reaction ID

A. Photolytic reactions

1) $O_2 + h\nu \rightarrow 2O$
2) $O_2 + h\nu \rightarrow O + O(^1D)$
3) $O_3 + h\nu \rightarrow O_2 + O$
4) $O_3 + h\nu \rightarrow O_2(^1\Delta) + O(^1D)$
5) $O_3 + h\nu \rightarrow O_2 + O(^1D)$
6) $O_3 + h\nu \rightarrow O_2(^1\Delta) + O$
7) $O_3 + h\nu \rightarrow 3O$
8) $H_2 + h\nu \rightarrow 2H$
9) $OH + h\nu \rightarrow O + H$
10) $HO_2 + h\nu \rightarrow OH + O$
11) $H_2O + h\nu \rightarrow H + OH$
12) $H_2O + h\nu \rightarrow H_2 + O(^1D)$
13) $H_2O + h\nu \rightarrow 2H + O$
14) $H_2O_2 + h\nu \rightarrow 2OH$
15) $N_2 + h\nu \rightarrow 2N$
16) $NO + h\nu \rightarrow N + O$
17) $NO_2 + h\nu \rightarrow NO + O$
18) $NO_3 + h\nu \rightarrow NO_2 + O$
19) $NO_3 + h\nu \rightarrow NO + O_2$
20) $N_2O + h\nu \rightarrow N_2 + O(^1D)$
21) $N_2O_5 + h\nu \rightarrow NO_2 + NO_3$
22) $N_2O_5 + h\nu \rightarrow NO + NO_3 + O$
23) $HNO_3 + h\nu \rightarrow NO_2 + OH$
24) $HO_2NO_2 + h\nu \rightarrow HO_2 + NO_2$
25) $HO_2NO_2 + h\nu \rightarrow OH + NO_3$
26) $CH_4 + h\nu \rightarrow CH_3 + H$
27) $CO_2 + h\nu \rightarrow CO + O$
28) $CO_2 + h\nu \rightarrow CO + O(^1D)$
29) $HCO + h\nu \rightarrow H + CO$
30) $H_2CO + h\nu \rightarrow 2H + CO$
31) $H_2CO + h\nu \rightarrow HCO + H$
32) $H_2CO + h\nu \rightarrow H_2 + CO$
33) $CH_3O_2 + h\nu \rightarrow CH_3 + O_2$
34) $CH_3OOH + h\nu \rightarrow CH_3O + OH$

B. Bi-/Ter-molecular reactions

52) $2O + M \rightarrow O_2 + M$
53) $2O + O_2 \rightarrow O_3 + O$
54) $O + 2O_2 \rightarrow O_3 + O_2$
55) $O + O_2 + N_2 \rightarrow O_3 + N_2$
56) $O + O_2 + CO \rightarrow O_3 + CO$
57) $O + O_2 + CO_2 \rightarrow O_3 + CO_2$
58) $O + O_2 + M \rightarrow O_3 + M$
59) $O + O_3 \rightarrow 2O_2$
60) $O + H + M \rightarrow OH + M$
61) $O + H_2 \rightarrow OH + H$
62) $O + OH \rightarrow O_2 + H$
63) $O + HO_2 \rightarrow OH + O_2$
64) $O + HO_2 \rightarrow OH + O_2(^1\Delta)$
65) $O + H_2O_2 \rightarrow OH + HO_2$
66) $O + NO + M \rightarrow NO_2 + M$
67) $O + NO_2 \rightarrow NO + O_2$
68) $O + NO_2 + M \rightarrow NO_3 + M$
69) $O + NO_3 \rightarrow O_2 + NO_2$
70) $O + N_2O_5 \rightarrow 2NO_2 + O_2$
71) $O + HNO_3 \rightarrow OH + NO_3$
72) $O + HO_2NO_2 \rightarrow OH + NO_2 + O_2$
73) $O + CH_3 \rightarrow H_2CO + H$
74) $O + CH_3 \rightarrow CO + H_2 + H$
75) $O + CH_4 \rightarrow CH_3 + OH$
76) $O + CO + M \rightarrow CO_2 + M$
77) $O + 2CO \rightarrow CO_2 + CO$
78) $2O + CO \rightarrow CO_2 + O$
79) $O + HCO \rightarrow H + CO_2$
80) $O + HCO \rightarrow OH + CO$
81) $O + H_2CO \rightarrow OH + HCO$
82) $O + CH_3O \rightarrow H_2CO + OH$
83) $O + CH_3O \rightarrow CH_3 + O_2$
84) $O + CH_3O_2 \rightarrow H_2CO + HO_2$
85) $O(^1D) + O_2 \rightarrow O + O_2$
86) $O(^1D) + N_2 \rightarrow O + N_2$
87) $O(^1D) + CO_2 \rightarrow O + CO_2$
88) $O(^1D) \rightarrow O$
89) $O(^1D) + O_3 \rightarrow 2O_2$
90) $O(^1D) + O_3 \rightarrow 2O + O_2$
91) $O(^1D) + H_2 \rightarrow H + OH$
92) $O(^1D) + H_2O \rightarrow 2OH$
93) $O(^1D) + N_2 + M \rightarrow N_2O + M$
94) $O(^1D) + N_2O \rightarrow 2NO$
95) $O(^1D) + N_2O \rightarrow N_2 + O_2$
96) $O(^1D) + CH_4 \rightarrow CH_3 + OH$
97) $O(^1D) + CH_4 \rightarrow CH_3O + H$
98) $O(^1D) + CH_4 \rightarrow H_2CO + H_2$
99) $O_2(^1\Delta) + O \rightarrow O_2 + O$
100) $O_2(^1\Delta) + O_2 \rightarrow 2O_2$
101) $O_2(^1\Delta) + H_2O \rightarrow O_2 + H_2O$
102) $O_2(^1\Delta) + N_2 \rightarrow O_2 + N_2$
103) $O_2(^1\Delta) + CO \rightarrow O_2 + CO$
104) $O_2(^1\Delta) + CO2 \rightarrow O_2 + CO_2$
105) $O_2(^1\Delta) + U \rightarrow O_2$
106) $O_2(^1\Delta) + O_3 \rightarrow 2O_2 + O$
107) $O_2(^1\Delta) + N \rightarrow NO + O$
108) $O_3 + NO \rightarrow NO_2 + O_2$
109) $O_3 + NO_2 \rightarrow NO_3 + O_2$
110) $H + O_2 + M \rightarrow HO_2 + M$
111) $H + O_3 \rightarrow OH + O_2$
112) $2H + M \rightarrow H_2 + M$
113) $H + OH + N_2 \rightarrow H_2O + N_2$
114) $H + OH + CO_2 \rightarrow H_2O + CO_2$
115) $H + HO_2 \rightarrow 2OH$
116) $H + HO_2 \rightarrow H_2 + O_2$
117) $H + HO_2 \rightarrow H_2 + O_2(^1\Delta)$
118) $H + HO_2 \rightarrow H_2O + O$
119) $H + NO_2 \rightarrow OH + NO$
120) $H + NO_3 \rightarrow OH + NO_2$
121) $H + CH_3 + M \rightarrow CH_4 + M$
122) $H + CH_4 \rightarrow CH_3 + H_2$
123) $H + CO + M \rightarrow HCO + M$
124) $H + HCO \rightarrow H_2 + CO$
125) $H + H_2CO \rightarrow H_2 + HCO$
126) $H + CH_3O \rightarrow H_2CO + H_2$
127) $H + CH_3O \rightarrow OH + CH_3$
128) $H + CH_3O_2 \rightarrow CH_4 + O_2$
129) $H + CH_3O_2 \rightarrow H_2O + H_2CO$
130) $OH + O_3 \rightarrow HO_2 + O_2$
131) $OH + O_3 \rightarrow HO_2 + O_2(^1\Delta)$
132) $OH + H_2 \rightarrow H_2O + H$
133) $2OH \rightarrow H_2O + O$
134) $2OH + M \rightarrow H_2O_2 + M$
135) $OH + HO_2 \rightarrow H_2O + O_2$
136) $OH + HO_2 \rightarrow H_2O + O_2(^1\Delta)$
137) $OH + H_2O_2 \rightarrow H_2O + HO_2$
138) $OH + NO_2 + M \rightarrow HNO_3 + M$
139) $OH + NO_3 \rightarrow HO_2 + NO_2$

140) $OH + HNO_3 \rightarrow NO_3 + H_2O$
141) $OH + HO_2NO_2 \rightarrow H_2O + NO_2 + O_2$
142) $OH + CH_3 \rightarrow H_2CO + H_2$
143) $OH + CH_3 \rightarrow CH_3O + H$
144) $OH + CH_3 \rightarrow CO + 2H_2$
145) $OH + CH_4 \rightarrow CH_3 + H_2O$
146) $OH + CO \rightarrow CO_2 + H$
147) $OH + HCO \rightarrow H_2O + CO$
148) $OH + H_2CO \rightarrow HCO + H_2O$
149) $OH + CH_3O \rightarrow H_2O + H_2CO$
150) $OH + CH_3OOH \rightarrow CH_3O_2 + H_2O$
151) $HO_2 + O_3 \rightarrow OH + 2O_2$
152) $2HO_2 \rightarrow H_2O_2 + O_2$
153) $2HO_2 \rightarrow H_2O_2 + O_2(^1\Delta)$
154) $2HO_2 + M \rightarrow H_2O_2 + O_2 + M$
155) $HO_2 + NO \rightarrow NO_2 + OH$
156) $HO_2 + NO_2 + M \rightarrow HO_2NO_2 + M$
157) $HO_2 + NO_3 \rightarrow HNO_3 + O_2$
158) $HO_2 + HCO \rightarrow H_2CO + O_2$
159) $HO_2 + CH_3O \rightarrow H_2CO + H_2O_2$
160) $HO_2 + CH_3O_2 \rightarrow CH_3OOH + O_2$
161) $N + O_2 \rightarrow NO + O$
162) $N + O_3 \rightarrow NO + O_2$
163) $N + OH \rightarrow NO + H$
164) $N + HO_2 \rightarrow NO + OH$
165) $2N + M \rightarrow N_2 + M$
166) $N + NO \rightarrow N_2 + O$
167) $N + NO_2 \rightarrow N_2O + O$
168) $N_2 \rightarrow 2N$
169) $NO + NO_3 \rightarrow 2NO_2$
170) $NO + CH_3O_2 \rightarrow CH_3O + NO_2$
171) $NO_3 + NO_2 \rightarrow NO + NO_2 + O_2$
172) $NO_3 + NO_2 + M \rightarrow N_2O_5 + M$
173) $2NO_3 \rightarrow 2NO_2 + O_2$
174) $N_2O_5 + H_2O \rightarrow 2HNO_3$
175) $N_2O_5 + M \rightarrow NO_3 + NO_2 + M$
176) $HO_2NO_2 + M \rightarrow HO_2 + NO_2 + M$
177) $CH_3 + O_2 \rightarrow H_2CO + OH$
178) $H_3 + O_2 + M \rightarrow CH_3O_2 + M$
179) $CH_3 + O_3 \rightarrow CH_3O + O_2$
180) $CH_3 + H_2 \rightarrow CH_4 + H$
181) $CH_3 + HO_2 \rightarrow CH_4 + O_2$
182) $CH_3 + HO_2 \rightarrow CH_3O + OH$
183) $CH_3 + H_2O_2 \rightarrow CH_4 + HO_2$
184) $CH_3 + HCO \rightarrow CH_4 + CO$
185) $CH_3 + H_2CO \rightarrow CH_4 + HCO$
186) $CH_3 + CH_3O \rightarrow H_2CO + CH_4$
187) $CO + NO_3 \rightarrow NO_2 + CO_2$
188) $HCO + O_2 \rightarrow CO + HO_2$
189) $2HCO \rightarrow H_2CO + CO$
190) $HCO + H_2CO \rightarrow CH_3O + CO$
191) $H_2CO + NO_3 \rightarrow HCO + HNO_3$
192) $CH_3O + O_2 \rightarrow H_2CO + HO_2$
193) $CH_3O + CO \rightarrow CH_3 + CO_2$
194) $CH_3O_2 + O_3 \rightarrow 2O_2 + CH_3O$
195) $2CH_3O_2 \rightarrow 2CH_3O + O_2$