

# Lawrence Berkeley National Laboratory

## LBL Publications

### Title

STRUCTURE OF TRIS(DIS(TRIMETHYLSILYL)AMIDO)NEODYMIUM (III), Nd[N(Si(CE3)3)2]3

### Permalink

<https://escholarship.org/uc/item/06x3n1xr>

### Author

Andersen, Richard A.

### Publication Date

1978-03-01

STRUCTURE OF TRIS(BIS(TRIMETHYLSILYL)AMIDO)NEODYMIUM(III),  
 $\text{Nd}[\text{N}(\text{Si}(\text{CH}_3)_3)_2]_3$

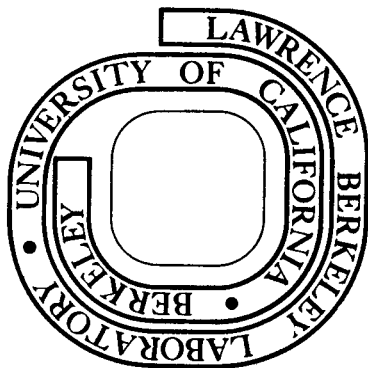
Richard A. Andersen, David H. Templeton, and  
Allan Zalkin

March 1978

Prepared for the U. S. Department of Energy  
under Contract W-7405-ENG-48

**TWO-WEEK LOAN COPY**

This is a Library Circulating Copy  
which may be borrowed for two weeks.  
For a personal retention copy, call  
Tech. Info. Division, Ext. ~~5716~~ 6782



Submitted to Inorganic Chemistry

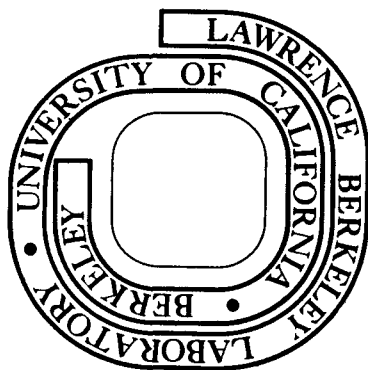
LBL-7603  
Preprint

STRUCTURE OF TRIS(BIS(TRIMETHYLSILYL)AMIDO)NEODYMIUM(III),  
 $\text{Nd}[\text{N}(\text{Si}(\text{CH}_3)_3)_2]_3$

Richard A. Andersen, David H. Templeton, and  
Allan Zalkin

March 1978

Prepared for the U. S. Department of Energy  
under Contract W-7405-ENG-48



LBL-7603

## **DISCLAIMER**

This document was prepared as an account of work sponsored by the United States Government. While this document is believed to contain correct information, neither the United States Government nor any agency thereof, nor the Regents of the University of California, nor any of their employees, makes any warranty, express or implied, or assumes any legal responsibility for the accuracy, completeness, or usefulness of any information, apparatus, product, or process disclosed, or represents that its use would not infringe privately owned rights. Reference herein to any specific commercial product, process, or service by its trade name, trademark, manufacturer, or otherwise, does not necessarily constitute or imply its endorsement, recommendation, or favoring by the United States Government or any agency thereof, or the Regents of the University of California. The views and opinions of authors expressed herein do not necessarily state or reflect those of the United States Government or any agency thereof or the Regents of the University of California.

Contribution from Materials and Molecular  
Research Division, Lawrence Berkeley Laboratory,  
and the Department of Chemistry,  
University of California, Berkeley, California 94720

Structure of Tris(bis(trimethylsilyl)amido)neodymium(III),  
 $\text{Nd}[\text{N}(\text{Si}(\text{CH}_3)_2)_3]_3^1$

By Richard A. Andersen\*, David H. Templeton\* and Allan Zalkin\*

Neodymium tris[di(trimethylsilyl)amide] is the only known monomeric, three coordinate derivative of this lanthanide element.<sup>2</sup> The structures of the europium (III)<sup>3</sup> and ytterbium (III)<sup>4</sup> derivatives have been examined by X-ray crystallographic techniques and they, along with the scandium (III) analogue,<sup>3</sup> have been shown to have  $\text{MN}_3$  skeletons which are not planar. In contrast all other crystallographically known tris-silylamides of the type  $\text{M}[\text{N}(\text{SiMe}_3)_2]_3$  are planar.<sup>4</sup> We describe the crystal structure of  $\text{Nd}[\text{N}(\text{SiMe}_3)_2]_3$  and show that it is also non-planar.

## EXPERIMENTAL

The  $\text{Nd}[\text{N}(\text{SiMe}_3)_2]_3$  was prepared as previously described,<sup>2</sup> m.p. 157-161° (lit.<sup>2</sup> 161-164°). The crystal used in the X-ray analysis was taken from a batch crystallized from pentane (0°C).

Magnetic susceptibility measurements were obtained with a PAR model 155 vibrating sample magnetometer employing a homogeneous magnetic field produced by a Varian Associates 12-inch electromagnet capable of a maximum field strength of 12.5 kG. The magnetometer was calibrated with  $\text{HgCo}(\text{CNS})_4$ .<sup>5</sup> A variable temperature liquid helium system produced sample temperatures in the range 4-100°K. The temperature was measured with a calibrated GaAs diode.

A hexagonal needle shaped crystal, 0.09 mm across and 0.3 mm long, was sealed inside a quartz capillary in an argon filled dry box. It was examined with a Picker FACS-I automatic diffractometer equipped with a graphite monochromator and a Mo x-ray tube ( $\lambda(\text{K}\alpha_1)$  0.7093 Å).  $\omega$  scans of several low-angle reflections showed peaks with half-widths of 0.16° and 0.21° for an h00 and 001 type reflections respectively. The space group was identified as  $P\bar{3}1c$ . The setting angles of 12 manually centered reflections ( $19^\circ < 2\theta < 25^\circ$ ) were used to determine by least squares the cell parameters  $a = 16.476(13)$  Å,  $c = 8.485(7)$  Å, and  $V = 1995$  Å<sup>3</sup>. For  $Z = 2$  and a molecular weight of 625.4 the calculated density is  $1.04 \text{ g cm}^{-3}$ .

Intensity data were collected using the  $\theta$ - $2\theta$  scan technique with a scan speed of 2°/min on  $2\theta$ . Each peak was scanned from 0.75° before

the  $K\alpha_1$  peak to  $0.75^\circ$  after the  $K\alpha_2$  peak, and backgrounds were counted for 10 s at each end of the scan range, offset by  $0.5^\circ$ . The needle direction of the crystal was approximately parallel to the  $\phi$  axis of the diffractometer. The temperature during data collection was  $21 \pm 1^\circ\text{C}$ . Three standard reflections, (300, 060 and 002), were measured after every 200th scan; no significant variation was observed in the intensities of the first two reflections, and a 5 percent decay in intensity was observed for the 002 reflection. A linear decay correction of about 4% was applied uniformly to the data. The absorption coefficient is estimated to be  $15 \text{ cm}^{-1}$ . Because of the diffraction geometry and the small crystal dimensions the absorption is small and no correction was deemed necessary. A total of 3404 scans, not including standards, resulted in 917 unique reflections, 535 of which were greater than  $2\sigma$ .

The positions of the Nd, N, and Si atoms were deduced from a three dimensional Patterson function. The carbon atoms were obtained from a subsequent least squares and Fourier calculation. A series of least-squares refinements in which the function  $\sum w(|F_o| - |F_c|)^2 / \sum w F_o^2$  was minimized converged to the final structure. The expressions that were used in processing the data and estimating the weights are given in the supplementary material; the "ignorance factor",  $p$ , was set to 0.06. Scattering factors from Doyle and Turner<sup>6</sup> were used, and dispersion corrections<sup>7</sup> were applied. Hydrogen atoms could not be located, and were not included. Because of the large residuals exhibited by several of

the low angle intensities, all 35 data whose  $(\sin\theta)/\lambda$  is less than 0.127, were deleted. The discrepancy index for 522 data are

$$R = \Sigma ||F_o| - |F_c|| / \Sigma |F_o| = 0.076$$

$$R_w = [\Sigma w(|F_o| - |F_c|)^2 / \Sigma w|F_o|^2]^{1/2} = 0.091$$

R for all 882 data is 0.13. The error in an observation of unit weight is 2.0. In the last cycle no parameter changed more than  $0.001\sigma$ .

#### RESULTS AND DISCUSSION

Atomic parameters, distances and angles are listed in Tables I-III. The molecular structure of this Nd complex (Fig. 1) is similar to that found in the Al,<sup>8</sup> Fe,<sup>9</sup> Sc,<sup>3</sup> and Eu<sup>3</sup> compounds. The Nd atom is on a crystallographic 3-fold axis and is bonded to three nitrogen atoms. In the Sc and Eu isomorphs<sup>4</sup> the metal atoms are disordered in the z direction, and were treated as two half atoms  $\sim 0.6 \text{ \AA}$  above and below the plane at  $z = 1/4$ . This structure is similar with Nd  $0.34 \text{ \AA}$  above and below  $z = 1/4$ . Originally the Nd atom was treated as an anisotropic atom at  $z = 1/4$  which resulted in thermal parameters  $B_{11}$  and  $B_{33}$  being 3.5 and  $16.0 \text{ \AA}^2$  respectively. When the Nd atom was treated isotropically as two half atoms disordered across the plane at  $z = 1/4$  the subsequent least-squares refinement resulted in the R factor going from 0.084 for the ordered to 0.076 for the disordered description.



A large channel that runs up the z axis, at the origin of the unit cell, is characteristic of the structures of these hexagonal  $M[N(SiMe_3)_2]_3$  complexes. Hursthouse and Rodesiler<sup>9</sup> have shown that in the case of the iron complex the channel is large enough to accommodate a benzene ring with the plane of the ring perpendicular to the z axis; they could not find any ordered solvent in the channel. A search of the final electron density and difference maps for the Nd structure showed one peak at 0,0,1/4 of about  $3 e/\text{\AA}^3$ , and three peaks between 0.4 and  $1.0 e/\text{\AA}^3$  just off the axis. The pattern of these peaks did not resemble any reasonable molecule that might have been used in the synthesis. It must be presumed that the channel, if occupied, contains solvent molecules that are so irregularly located as to be virtually invisible to the x-ray diffraction technique. The large R factor, and the large error of a reflection of unit weight may be a result of this unresolved structure.

The variable temperature (4.2 - 89.6°K) magnetic susceptibility follows Curie-Weiss behavior,  $\chi = \frac{C_M}{T+\theta}$ ,  $C_M = 1.33$  and  $\theta = 12^\circ\text{K}$ . The magnetic moment,  $\mu_{\text{eff}}$ , is 3.27 B.M.

Supplementary Material Available: Data processing formulas and the listing of structure factor amplitudes (5 pages). Ordering information is given on any current masthead.

#### ACKNOWLEDGEMENT

We thank Dr. N. M. Edelstein for useful discussions.

REFERENCES

1. This work was supported by the Division of Basic Energy Sciences of the Department of Energy.
2. D. C. Bradley, J. S. Ghotra, and F. A. Hart, *J.C.S. Dalton*, 1021 (1973).
3. J. S. Ghotra, M. B. Hursthouse, and A. J. Welch, *Chem. Commun.* 669 (1973).
4. P. G. Eller, D. C. Bradley, M. B. Hursthouse, and D. W. Meek, *Coord. Chem. Rev.* (24) 1 (1977).
5. H. St. Råde, *J. Phys. Chem.* 77, 424 (1973).
6. P. A. Doyle and P. S. Turner, *Acta Crystallogr. Sect A*, 24, 390 (1968).
7. D. T. Cromer and D. Liberman, *J. Chem. Phys.* 53, 1891 (1970).
8. G. M. Sheldrick and W. S. Sheldrick, *J. Chem. Soc. A*, 2279 (1969).
9. M. B. Hursthouse and P. F. Rodesiler, *J. Chem. Soc. Dalton*, 2100 (1972).

Table I. Positional and Thermal Parameters<sup>a</sup> with Estimated Deviations<sup>b</sup> for Nd[N(Si(CH<sub>3</sub>)<sub>3</sub>)<sub>2</sub>]<sub>3</sub>

ATOM	x	y	z	B <sub>11</sub>	B <sub>22</sub>	B <sub>33</sub>	B <sub>12</sub>	B <sub>13</sub>	B <sub>23</sub>
Nd <sup>c</sup>	2/3	1/3	.2903(3)	3.49(7)					
N <sup>c</sup>	.508(1)	.254	1/4	4.3(7)	4.7(6)	2.4(7)	2.15	0	.3(5)
Si	.4543(3)	.2866(3)	.1131(5)	4.5(2)	5.4(2)	4.0(2)	2.8(2)	-.6(2)	-.1(2)
C(1)	.401(1)	.353(1)	.202(2)	8(1)	9(1)	8(1)	6(1)	1(1)	-1(1)
C(2)	.364(1)	.187(1)	-.005(2)	6(1)	8(1)	7(1)	2(1)	-4(1)	-1(1)
C(3)	.546(1)	.367(1)	-.033(2)	7(1)	8(1)	4(1)	4(1)	1(1)	3(1)

<sup>a</sup>The anisotropic temperature factor has the form  $\exp(-0.25(B_{11}h^2a^{*2} + 2B_{12}hka^*b^* + \dots))$ .

<sup>b</sup>Here and in the following tables the numbers in parenthesis are the estimated standard deviations in the least significant digit.

<sup>c</sup>Symmetry conditions of the special positions for N;  $x = 2y$ ,  $B_{11} = 2B_{12}$ , and  $B_{13} = 0$ .

Table II. Interatomic Distances (Å)

		Corrected <sup>a</sup>
Nd - 3N	2.29(2)	2.29
N - 2Si	1.70(1)	1.71
Si - C(1)	1.88(2)	1.91
Si - C(2)	1.86(2)	1.90
Si - C(3)	1.89(2)	1.91

<sup>a</sup>Adjusted for thermal motion assuming the "riding" model.

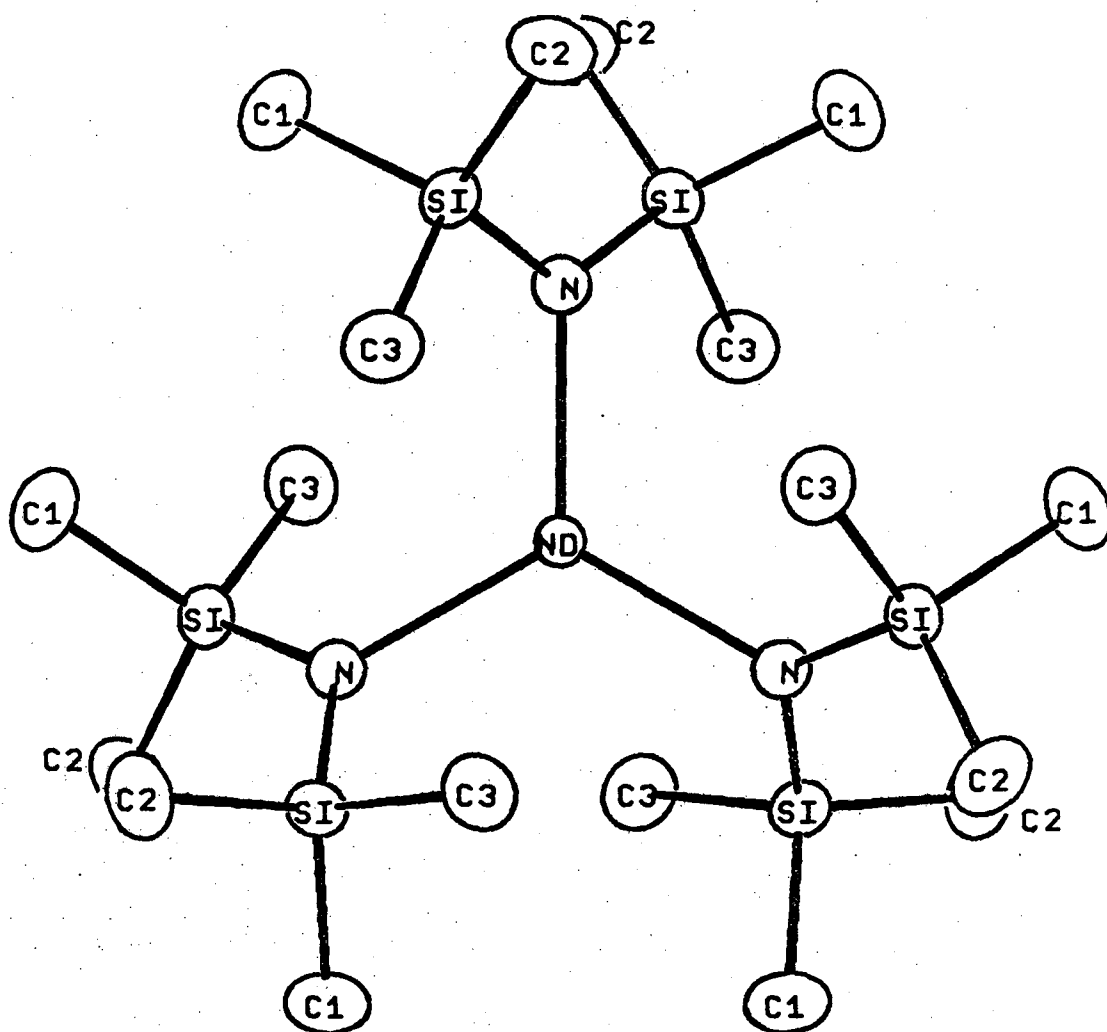
Table III. Selected Angles (deg.)

N	- Nd - N	117.8(1)
Nd	- N - Si	123.2(5)
Nd	- N - Si <sup>a</sup>	110.1(4)
Si	- N - Si <sup>a</sup>	126.4(9)
N	- Si - C(1)	112.4(6)
N	- Si - C(2)	114.0(7)
N	- Si - C(3)	108.0(7)
C(1)	- Si - C(2)	108.9(9)
C(1)	- Si - C(3)	107.2(8)
C(2)	- Si - C(3)	105.9(8)

<sup>a</sup>Atom at position x, x-y, 1/2 - z.

FIGURE CAPTION

Fig. 1. ORTEP view of  $\text{Nd}(\text{N}(\text{SiMe}_3)_2)_3$  down the c axis.

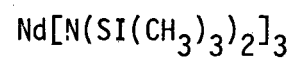


XBL 782-7162

Fig. 1

SUPPLEMENTARY MATERIALS FOR THE PAPER

Structure of Tris(bis(trimethylsily)amido)neodymium(III),



by Richard A. Andersen\*, David H. Templeton\* and Allan Zalkin\*



DATA PROCESSING FORMULAE

$$I = C - (t_c/2t_b)(B_1+B_2)$$

$$\sigma(B) = \text{Max}[(t_c/2t_b)(B_1+B_2)^{\frac{1}{2}}, (t_c/2t_b)|B_1-B_2|]$$

$$\sigma(I) = [0 + \sigma^2(B)]^{\frac{1}{2}}$$

$$F^2 = (D \cdot A / L_p) I$$

$$\sigma(F^2) = (D \cdot A / L_p) \sigma(I)$$

$$F_a^2 = \Sigma F^2 / n$$

$$\sigma(F_a^2) = [\Sigma \sigma^2(F^2) / n]^{\frac{1}{2}}$$

When  $S(F_a^2) > 4\sigma(F_a^2)$ ,  $\sigma(F_a^2)$  is replaced by  $S(F_a^2)$ .

$$S(F_a^2) = [\Sigma |F^2 - F_a^2|^2 / n(n-1)]^{\frac{1}{2}}$$

$$\sigma(F_o^2) = [\sigma^2(F_a^2) + (pF_a^2)^2 + q^2]^{\frac{1}{2}}$$

$$F_o = (F_a^2)^{\frac{1}{2}}$$

$$\sigma(F) = F_o - [F_a^2 - \sigma(F_o^2)]^{\frac{1}{2}} \text{ when } \sigma(F_o^2) \leq F_a^2 \text{ or } [\sigma(F_a^2)]^{\frac{1}{2}} \text{ when } \sigma(F_a^2) > F_a^2$$

$$L_p = [\cos^2 2\theta_m + \cos^2 2\theta] / [\sin 2\theta (1 + \cos^2 2\theta_m)]$$

$$\text{wtg} = 1/\sigma^2(F)$$

C = counts recorded during a scan

$\theta_m$  = monochromater angle

I = individual raw intensity,  
background removed.

$\theta$  = crystal diffraction angle

$t_c$  = scan count time

S = scatter

$t_b$  = background count time

a = average

$B_1$  = individual background count

q = additional uncertainty that  
affects the weak intensities

$\sigma(B)$  = estimated standard deviation of the total background count

p = estimate of non-statistical errors

F = structure factor

wtg = weighting factors in least squares

D = decay correction; an empirically applied correction obtained from the fluctuations of the standard reflections.

A = absorption correction

$L_p$  = Lorentz and polarization corrections

OBSERVED STRUCTURE FACTORS, STANDARD DEVIATIONS, AND DIFFERENCES (ALL X 6.0)  
F(0,0,0) = 3276

FOB AND FCA ARE THE OBSERVED AND CALCULATED STRUCTURE FACTORS.  
SG = ESTIMATED STANDARD DEVIATION OF FOB. DEL = /FOB/ - /FCA/.  
\* INDICATES ZERO WEIGHTED DATA.

L	FOB	SG	DEL	L	FOB	SG	DEL	L	FOB	SG	DEL	L	FOB	SG	DEL	L	FOB	SG	DEL
H,K= 0, 0	6	127	10	-32	8	25	55	6*	2	388	12	27	H,K= 5, 3						
2 406 0 131*	8	193	11	3	H,K= 4, -1	4	431	14	-19	0	39	11	19						
4 208 7 22	H,K= 3, -1	1	591	19	64	6	33	80	21*	1	493	15	19						
6 94 12 -37	1	860	0	38*	2	668	21	22	8	0	63	-52*	2	136	6	-16			
8 114 12 -9	2	461	15	-40	3	373	12	10	H,K= 5, -2	3	252	8	-12						
H,K= 1, 0	3	557	17	11	4	78	9	12	1	678	21	9	4	347	11	2			
11048 0 12*	4	256	8	-18	5	17	36	-14*	2	168	8	54	5	157	8	-3			
2 102 0 -147*	5	40	20	24*	6	51	26	-18*	3	139	6	-9	6	72	53	2*			
3 334 11 45	6	80	11	-37	7	133	21	-25	4	117	8	12	7	26	43	14*			
4 62 8 -34	7	116	21	-12	8	42	48	8*	5	54	24	30*	8	56	35	-1*			
5 43 13 36*	8	65	15	-19	H,K= 4, 0	6	41	42	31*	H,K= 5, 4									
6 105 8 47	9	26	56	19*	0	313	13	26	7	48	72	-16*	0	289	9	-50			
7 172 10 2	H,K= 3, 0	1	851	26	46	H,K= 5, -1	1	394	13	46	1	176	7	-10					
8 38 31 32*	0	482	0	186*	2	32	42	0*	2	459	14	35	2	51	12	34			
9 69 20 -5	1	404	0	35*	3	179	6	-11	3	209	7	15	3	233	9	-13			
H,K= 1, 1	2	899	27	140	4	169	6	-8	4	34	34	-20*	4	181	9	-22			
0 914 0 306*	3	180	7	-7	5	65	10	-33	5	78	19	0	5	200	8	1			
2 231 15 93	4	299	9	-4	6	94	8	-31	6	33	43	-8*	6	103	22	1			
4 226 9 27	5	53	15	37	7	27	67	7*	7	0	63	-36*	7	51	39	26*			
6 85 28 -42*	6	58	12	23	8	76	14	-11	8	34	57	8*	H,K= 5, 5						
8 103 22 -19	7	144	13	2	H,K= 4, 1	0	560	17	-21	H,K= 5, 0	0	382	12	44					
H,K= 2, -1	8	139	12	-18	1	21	59	-42*	0	28	47	-99*	2	213	8	45			
2 986 30 146	H,K= 3, 1	1	21	59	-42*	2	508	16	24	1	529	16	11	4	194	11	-27		
4 352 11 20	0	331	0	-100*	3	85	7	0	3	45	16	29	6	51	63	7*			
6 67 20 63*	1	142	23	12	4	95	7	19	2	403	12	20	H,K= 6, -3						
8 63 30 41*	2	298	11	67	5	281	10	-5	4	38	16	2*	2	264	10	-6			
H,K= 2, 0	3	377	12	18	6	48	23	-9*	5	38	39	-11*	4	223	9	4			
0 373 0 -266*	4	69	8	-18	7	25	69	-9*	6	123	8	22	6	65	28	-12*			
1 297 0 -8*	5	160	7	6	8	39	51	-21*	7	39	49	-1*	8	58	51	-11*			
2 75 47 -37*	6	232	8	22	H,K= 4, 2	0	277	9	0	8	35	45	27*	H,K= 6, -2					
3 508 16 16	7	113	22	-29	0	277	9	0	0	366	12	-38	1	341	11	26			
4 92 5 -16	8	54	64	6*	1	638	20	-2	H,K= 5, 1	1	362	11	-35	2	481	15	14		
5 138 6 -1	H,K= 3, 2	0	118	20	-91	2	13	33	-15*	0	406	13	1	3	95	7	4		
6 101 8 60	1	166	12	-46	3	284	9	3	1	362	11	-35	4	192	7	-14			
7 93 19 -8	2	259	9	7	4	246	8	10	2	406	13	1	5	237	9	-4			
8 48 20 32*	3	281	9	4	5	195	9	2	3	49	15	-21*	6	76	14	-2			
9 62 26 -24*	4	41	16	19*	6	99	11	38	4	237	8	-2	7	70	29	-3*			
H,K= 2, 1	5	205	9	2	7	6	61	-6*	5	59	15	1	8	43	48	33*			
0 449 0 219*	6	26	45	-14*	8	43	51	-29*	6	96	19	-8	H,K= 6, -1						
1 155 0 58*	7	26	73	12*	H,K= 4, 3	0	283	9	-33	7	18	59	-7*	1	389	12	-35		
2 51 71 -75*	8	0	62	-29*	0	283	9	-33	H,K= 5, 2	8	79	20	49	2	262	8	60		
3 206 8 -22	H,K= 3, 3	0	612	19	53	1	558	17	36	0	529	16	60	3	221	7	-17		
4 215 7 2	0	612	19	53	2	521	16	-8	0	529	16	60	4	26	36	10*			
5 284 10	2	800	24	81	3	452	14	-3	1	243	8	2	5	85	12	1			
6 47 35 40*	4	105	11	1	4	33	48	-8*	2	605	19	36	6	203	8	23			
7 77 44 1*	6	53	39	45*	5	86	11	-11	3	46	23	13*	7	24	47	0*			
8 115 9 -43	8	59	48	19*	6	53	55	48*	4	116	8	-18	8	40	62	-2*			
9 110 15 -24	H,K= 4, -2	2	562	16	-68	7	76	19	-5	5	194	8	-2	H,K= 6, 0					
H,K= 2, 2	2	562	16	-68	8	46	51	33*	6	87	22	37	0	817	25	86			
0 531 0 159*	4	296	10	-4	H,K= 4, 4	0	600	18	72	7	81	11	12	1	25	35	-12*		
2 0 69 -26*	6	305	11	-19	0	600	18	72	8	55	67	19*	2	184	6	2			
4 445 14 7														3	37	15	28*		



STRUCTURE FACTORS CONTINUED FOR

L	FOB	SG	DEL	L	FOB	SG	DEL	L	FOB	SG	DEL	L	FOB	SG	DEL	L	FOB	SG	DEL
7	69	21	41*	H,K= 9,	7	2	374	12	25	6	57	28	9*	1	96	9	-7		
	H,K= 9,	0		0	111	9	-10	3	91	12	20	7	45	50	15*	2	23	48	-1*
0	144	7	21	1	220	9	9	4	114	22	-9		H,K= 11,	-3	3	148	10	15	
1	31	32	7*	2	109	18	27	5	124	8	6	1	354	11	-2	4	52	52	9*
2	302	10	24	3	42	49	9*	6	60	31	33*	2	130	9	-11		H,K= 11,	5	
3	171	7	7		H,K= 9,	8			H,K= 10,	2		3	49	21	35*	0	151	8	11
4	155	8	-9	6	99	10	-3	0	101	10	-6	4	50	60	-2*	1	76	20	18
5	57	59	10*	1	197	13	5	1	330	11	8	5	0	56	-5*	2	152	11	25
6	43	19	29*	2	62	34	2*	2	85	16	1	6	52	37	44*	3	62	35	27*
7	60	23	29*		H,K= 10,	-5		3	162	19	1	7	40	51	9*		H,K= 11,	6	
	H,K= 9,	1		2	509	17	18	4	94	22	-12		H,K= 11,	-2	0	44	28	22*	
0	26	37	-42*	4	151	12	6	5	66	26	35*	1	318	10	11	1	147	10	2
1	346	11	3	6	23	55	20*	6	29	69	-23*	2	122	9	23	2	81	19	-6
2	60	14	17		H,K= 10,	-4			H,K= 10,	3		3	113	10	1		H,K= 12,	-6	
3	163	7	-2	1	572	18	11	0	66	23	14*	4	39	68	26*	2	0	47	-19*
4	59	36	-3*	2	0	43	-41*	1	119	12	4	5	6	47	-4*	4	215	12	-1
5	10	69	6*	3	251	9	4	2	100	25	-20	6	99	14	24	6	54	69	-8*
6	26	50	16*	4	120	10	-14	3	80	18	-24		H,K= 11,	-1		H,K= 12,	-5		
7	57	41	1*	5	63	66	14*	4	60	20	8*	1	32	42	30*	1	86	9	6
	H,K= 9,	2		6	37	53	-12*	5	30	55	-47*	2	201	9	25	2	41	35	24*
0	261	8	-18	7	66	22	64*		H,K= 10,	4		3	134	12	16	3	107	11	10
1	273	9	-4		H,K= 10,	-3		0	164	11	12	4	134	18	-10	4	79	41	-7*
2	210	8	-2	1	345	11	-3	1	36	46	30*	5	21	52	6*	5	132	9	-12
3	100	12	12	2	298	10	-30	2	0	59	-56*	6	26	51	-9*	6	29	50	-7*
4	60	70	25*	3	139	8	0	3	32	43	18*		H,K= 11,	0		H,K= 12,	-4		
5	49	23	0*	4	116	9	10	4	166	12	2	0	126	8	11	1	314	10	1
6	32	51	-31*	5	29	72	-7*	5	53	49	38*	1	265	9	2	2	59	17	-5
	H,K= 9,	3		6	56	21	13*		H,K= 10,	5		2	227	8	-8	3	98	14	-4
0	360	11	17	7	53	44	30*	0	43	32	-3*	3	153	10	-13	4	48	65	0*
1	97	9	0		H,K= 10,	-2		1	51	33	-25*	4	137	8	-7	5	33	44	17*
2	180	8	23	1	24	42	16*	2	69	13	3	5	64	18	7	6	10	50	3*
3	89	24	39	2	260	9	23	3	75	19	-12	6	25	50	10*		H,K= 12,	-3	
4	52	64	-7*	3	48	23	21*	4	119	13	-5		H,K= 11,	1		1	145	8	-7
5	61	17	-14	4	53	20	16*		H,K= 10,	6		0	169	7	1	2	194	8	18
6	73	21	37	5	46	63	29*	0	117	8	-2	1	256	9	7	3	71	32	6*
	H,K= 9,	4		6	51	31	4*	1	131	11	2	2	141	15	-3	4	0	68	-43*
0	135	7	2	7	33	61	24*	2	77	17	25	3	223	12	0	5	48	27	-17*
1	71	20	7		H,K= 10,	-1		3	139	11	13	4	27	49	17*	6	45	55	42*
2	53	71	43*	1	211	8	6		H,K= 10,	7		5	65	24	-8*		H,K= 12,	-2	
3	107	15	-12	2	111	8	-20	0	144	8	-4	6	55	37	45*	1	118	9	-5
4	30	47	6*	3	137	8	4	1	0	56	-10*		H,K= 11,	2		2	185	8	-12
5	81	20	11	4	66	32	15*	2	169	12	4	0	233	9	19	3	71	40	9*
6	101	16	-7	5	95	33	-2*		H,K= 11,	-5		1	33	62	-24*	4	26	47	-17*
	H,K= 9,	5		6	0	49	-20*	1	213	8	8	2	263	11	-3	5	93	15	15
0	42	54	3*	7	34	50	29*	2	185	8	-30	3	0	48	-5*	6	60	33	20*
1	52	71	-4*		H,K= 10,	0		3	104	10	1	4	62	44	-16*		H,K= 12,	-1	
2	0	61	-7*	0	299	10	-15	4	65	19	44*	5	0	67	-19*	1	129	8	-11
3	82	14	-31	1	153	6	6	5	120	15	-17		H,K= 11,	3		2	56	61	29*
4	85	36	-4*	2	119	6	-10	6	50	39	25*	0	134	12	-20	3	152	13	-10
5	115	13	2	3	238	8	7	7	18	51	-17*	1	123	15	-6	4	67	14	8
	H,K= 9,	6		4	33	58	-2*		H,K= 11,	-4		2	105	11	2	5	129	17	13
0	226	8	6	5	95	11	5	1	60	19	-8*	3	105	12	-3	6	41	52	-21*
1	27	41	14*	6	52	55	37*	2	294	9	13	4	53	37	20*		H,K= 12,	0	
2	105	11	-9		H,K= 10,	1		3	0	44	-5*	5	85	19	20	0	205	10	25
3	36	62	3*	0	285	9	17	4	84	15	56		H,K= 11,	4		1	37	52	32*
4	114	12	18	1	184	7	6	5	31	62	-6*	0	130	7	-3	2	236	10	-2

STRUCTURE FACTORS CONTINUED FOR

L	FOB	SG	DEL	L	FOB	SG	DEL	L	FOB	SG	DEL	L	FOB	SG	DEL	L	FOB	SG	DEL
3	19	45	6*	5	32	70	24*	5	0	75	-8*	1	181	10	5				
4	143	7	-3		H,K=	13,	-2		H,K=	14,	-3	2	60	61	49*				
5	46	58	31*	1	42	56	9*	1	193	11	5	3	86	16	8				
	H,K=	12,	1	2	139	17	-9	2	45	32	-1*		H,K=	15,	-1				
0	149	11	-14	3	32	41	-1*	3	156	8	1	1	91	16	-20				
1	191	9	-1	4	99	19	-28	4	71	21	9*	2	109	13	0				
2	158	10	-10	5	32	51	20*	5	0	52	-19*	3	86	25	5				
3	157	8	-6		H,K=	13,	-1		H,K=	14,	-2		H,K=	15,	0				
4	49	51	45*	1	134	17	-12	1	103	9	-21	0	106	14	10				
5	31	51	-16*	2	21	53	13*	2	161	10	11	1	0	47	-32*				
	H,K=	12,	2	3	147	8	0	3	50	54	10*	2	125	10	-8				
0	33	37	-20*	4	68	37	22*	4	64	24	8*		H,K=	16,	-8				
1	314	10	14	5	71	23	-20*		H,K=	14,	-1	2	151	13	6				
2	0	49	-43*		H,K=	13,	0	1	72	18	29		H,K=	16,	-7				
3	128	11	10	0	51	25	-9*	2	121	11	5	1	202	10	-3				
4	95	15	9	1	190	8	-9	3	56	61	-28*	2	0	52	-36*				
	H,K=	12,	3	2	119	8	7	4	84	27	4*	3	60	26	5*				
0	235	8	5	3	83	14	-4		H,K=	14,	0		H,K=	16,	-6				
1	48	42	13*	4	108	13	1	0	119	15	-9	1	102	13	-3				
2	145	13	-1		H,K=	13,	1	1	161	8	10	2	58	35	-27*				
3	0	50	-12*	0	241	8	4	2	63	21	-4*	3	106	17	-7				
4	31	51	-22*	1	105	10	-8	3	128	10	10		H,K=	16,	-5				
	H,K=	12,	4	2	209	10	7		H,K=	14,	1	1	75	18	30				
0	11	43	-9*	3	43	50	32*	0	95	13	17	2	85	17	-14				
1	197	11	22	4	35	51	-32*	1	164	10	5	3	39	55	-21*				
2	10	56	-18*		H,K=	13,	2	2	41	71	5*		H,K=	16,	-4				
3	82	20	-1	0	127	8	1	3	116	14	5	1	113	13	-17				
	H,K=	12,	5	1	168	10	16		H,K=	14,	2	2	21	49	-39*				
0	106	14	-5	2	144	10	5	0	150	10	0	3	0	74	-37*				
1	81	18	-15	3	84	17	8	1	0	63	-5*		H,K=	16,	-3				
	H,K=	13,	-6		H,K=	13,	3		H,K=	15,	-7	1	109	13	-15				
1	85	11	1	0	122	11	-3	1	258	10	12	2	54	38	5*				
2	41	52	25*	1	152	10	11	2	111	12	-9		H,K=	16,	-2				
3	113	21	-9	2	104	22	23	3	0	49	-59*	1	70	24	54*				
4	114	13	-4		H,K=	13,	4	4	34	50	3*		H,K=	17,	-8				
5	150	10	9	0	168	9	10		H,K=	15,	-6	1	164	9	20				
6	75	35	66*		H,K=	14,	-7	1	25	49	-10*	2	46	62	-26*				
	H,K=	13,	-5	2	152	18	8	2	232	8	5		H,K=	17,	-7				
1	127	8	3	4	126	16	10	3	0	49	-6*	1	34	68	31*				
2	205	9	9		H,K=	14,	-6	4	0	51	-44*	2	86	17	-16				
3	24	68	-5*	1	181	13	9		H,K=	15,	-5		H,K=	17,	-6				
4	154	9	-12	2	44	64	-55*	1	204	8	-2	1	57	28	-4*				
5	42	57	40*	3	155	8	-8	2	79	23	-8	2	84	27	24*				
6	78	24	46*	4	106	21	26	3	154	11	-5		H,K=	17,	-5				
	H,K=	13,	-4	5	43	55	36*	4	92	15	-6	1	78	18	4				
1	185	8	11		H,K=	14,	-5		H,K=	15,	-4								
2	191	12	-8	1	248	12	9	1	134	8	2								
3	186	11	-11	2	140	12	-16	2	88	15	-10								
4	20	55	11*	3	208	8	2	3	133	11	-13								
5	37	50	28*	4	47	58	-2*	4	45	63	23*								
6	12	57	-10*	5	34	50	26*		H,K=	15,	-3								
	H,K=	13,	-3		H,K=	14,	-4	1	73	17	13								
1	274	10	9	1	0	63	-18*	2	116	15	11								
2	79	34	7*	2	296	10	6	3	22	49	1*								
3	134	11	-11	3	50	20	29*	4	70	22	18*								
4	66	15	-9	4	134	11	11		H,K=	15,	-2								

This report was done with support from the Department of Energy. Any conclusions or opinions expressed in this report represent solely those of the author(s) and not necessarily those of The Regents of the University of California, the Lawrence Berkeley Laboratory or the Department of Energy.

TECHNICAL INFORMATION DEPARTMENT  
LAWRENCE BERKELEY LABORATORY  
UNIVERSITY OF CALIFORNIA  
BERKELEY, CALIFORNIA 94720